CHEMISTRY-11	Chapter#08 (Complete-Smart Syllabus) Test-1B		
	Name:	Class:	ID:
Date: / /	**Marks Total: 25**	**Marks Obtained:**	
Time Allowed: 40 Min.			

Maximum Marks: 09 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1. **For which system does the equilibrium constant, K_c, has units of (conc.)$^{-1}$?**
   (a) $N_2 + 3H_2 \rightleftharpoons 2NH_3$ (b) $H_2 + I_2 \rightleftharpoons 2HI$
   (c) $2NO_2 \rightleftharpoons N_2O_4$ (d) $2HF \rightleftharpoons H_2 + F_2$
2. **An excess of aqueous silver nitrate is added to aqueous barium chloride and precipitate is removed by filtration. What are the main ions in the filtrate?**
   (a) Ag^+ and NO_3^{-1} only (b) Ag^+ and Ba^{2+} and NO_3^{-1}
   (c) Ba^{2+} and NO_3^{-1} only (d) Ba^{2+} and NO_3^{-1} and Cl^{-1}
3. **For the reaction $2SO_2 + O_2 \rightleftharpoons 2SO_3$:**
   (a) $K_c = K_p$ (b) $K_c < K_p$ (c) $K_c > K_p$ (d) None
4. **The term pH was introduced by:**
   (a) Henderson (b) Sorenson (c) Goldstein (d) Thomson
5. **When HCl is added to aqueous solution of H_2S, its ionization:**
   (a) Increases (b) Remains constant
   (c) Decreases (d) First increases then decreases
6. **Change in concentration will change:**
   (a) Position of equilibrium (b) K_c
   (c) Both 'a' & 'b' (d) None
7. **The catalyst used in Haber's process is:**
   (a) Fe (b) Al_2O_3 (c) MgO (d) All
8. **The pH of milk is:**
   (a) 6.5 (b) 7.8 (c) 6.5 to 6.9 (d) 8.5
9. **Common ion effect is the application of:**
   (a) Law of mass action (b) Le-Chatelier's Principle
   (c) Buffer (d) All

Maximum Marks: 16 **(SUBJECTIVE TYPE)** Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: **(12)**

i. Differentiate between reversible and irreversible reactions.
ii. How does K_c predict about the direction of the chemical reaction?
iii. What is the effect of catalyst on a system at equilibrium?
iv. Define common ion effect?
v. Why do we need buffers?
vi. Define buffer capacity.

SECTION-II

NOTE: Attempt All Questions: **(04)**

Q.3: What is the percentage ionization of acetic acid in a solution in which 0.1 moles of it has been dissolved per dm^3 of the solution?