

CHEMISTRY-11	Chapter#11(Complete-Smart Syllabus) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Specific rate constant is equal to rate of reaction, when concentration of reactants is:
 - Zero
 - Four
 - Three
 - Unity
- The chemical weathering of stone work of buildings is caused by the _____ gases in the atmosphere.
 - Acidic
 - Basic
 - Radioactive
 - All
- The rate of the reaction $A \rightarrow B$, can be expressed as:
 - $\frac{d[A]}{dt}$
 - $\frac{-d[B]}{dt}$
 - $\frac{-d[A]}{dt}$
 - $\frac{-d[A]}{-dt}$
- The units of rate constant for zero order reaction are:
 - Moles $dm^{-3}s^{-1}$
 - Moles $^{-1}dm^3s^{-1}$
 - Moles $^{-2}dm^6s^{-1}$
 - s^{-1}
- Order of reaction can never be:
 - Zero
 - Negative
 - In fraction
 - More than three
- For a zero-order reaction, if the concentration of the reactants is doubled, the half-life:
 - Also becomes double
 - Becomes half
 - Becomes four times
 - Remains same
- The reaction $NO_2 + CO \rightarrow NO + CO_2$ is _____ order with respect to NO_2 .
 - First
 - Second
 - Third
 - Zero
- In the method of large excess, the rate and order of the reaction is controlled by that reactant which is taken in:
 - Small amount
 - Excess amount
 - Both 'a' & 'b'
 - None
- In sunlight, the reaction between H_2 and Cl_2 is:
 - Slow
 - Moderate
 - Explosive
 - Negligible
- The rate of reaction generally doubles with every _____ increase in temperature.
 - 10 K
 - 30 K
 - 20 K
 - 40 K

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(12)

- What is specific rate constant or velocity constant?
- What are zero order reactions? Give one example.
- The slowest step is the rate determining step. Explain.
- Name different methods for determining the order of reaction.
- Define half-life period. How is it used to determine the order of reaction?
- How does increase in temperature increases the rate of reaction?

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: How light and surface area affect the rate of chemical reaction.

Q.4: How does Arrhenius equation help us to calculate the activation energy of a reaction.