

<b>CHEMISTRY-11</b>	<b>Chapter#10 (Complete - Smart Syllabus) Test-4</b>		
	Name:	Class:	ID:
Date: / /	<b>Marks Total: 30</b>	<b>Marks Obtained:</b>	
Time Allowed: 50 Min.			

Maximum Marks: 10

**(OBJECTIVE TYPE)**

Time Allowed: 10 Min.

**NOTE:** Tick The Correct Option:

- If a strip of Cu metal is placed in a solution of FeSO<sub>4</sub>:
  - Cu will be deposited
  - Fe is precipitated out
  - Cu and Fe both dissolve
  - No reaction takes place
- When one metal is deposited on the surface of the other by the process of electrical current, it is called:
  - Electrolysis
  - Electroplating
  - Electrolytic refining
  - Electrolytic
- An oxidizing agent:
  - Is oxidized
  - Gains electrons
  - Neither gains nor loses electrons
  - Loses electrons
- In per oxides, the oxidation number of O is:
  - 0
  - $\frac{1}{2}$
  - 1
  - 2
- During the electrolysis of aqueous NaCl solution, \_\_\_\_\_ is deposited at anode.
  - Na
  - Cl<sub>2</sub>
  - H<sub>2</sub>
  - O<sub>2</sub>
- Reduction occurs at:
  - Anode
  - Cathode
  - Both 'a' & 'b'
  - None
- Castner-Kellner cell is used to obtain:
  - Sodium
  - Aluminum
  - Caustic soda
  - Copper
- In Galvanic cell, anode is \_\_\_\_\_ charged.
  - Positively
  - Negatively
  - Neutrally
  - Both 'a' & 'b'
- The voltage produced in Galvanic cell is:
  - 1.1 V
  - 1.4 V
  - 1.2 V
  - 1.5 V
- Which metal can release H<sub>2</sub> from steam?
  - Fe
  - Cu
  - Ag
  - Au

Maximum Marks: 20

**(SUBJECTIVE TYPE)**

Time Allowed: 40 Min.

**SECTION-I**

- Q.2: Give brief answers to the following questions: (12)**
- Calculate the oxidation number of Cr in K<sub>2</sub>CrO<sub>4</sub> and CrCl<sub>3</sub>.
  - Differentiate between oxidation and reduction.
  - How can copper be purified electrolytically?
  - Differentiate between single electrode potential and standard electrode potential.
  - What is electrochemical series?
  - Na and K can displace hydrogen from acids but Pt, Pd and Cu cannot. Why?

**SECTION-II**

**NOTE:** Attempt All Questions:

(08)

- Q.3: Define electrolysis? Explain the electrolysis of very dilute solution of NaNO<sub>3</sub>.**
- Q.4: Describe the construction and working of standard hydrogen electrode.**