

CHEMISTRY-11	Chapter#10 (Complete - Smart Syllabus) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The cathode reaction in the electrolysis of dil. H_2SO_4 with Pt electrodes is:
 - Reduction
 - Oxidation
 - Both oxidation and reduction
 - Neither oxidation or reduction
- In H_2O_2 , the oxidation state of oxygen is:
 - +1
 - 1
 - +2
 - 2
- Standard hydrogen electrode (SHE) is made of:
 - Ag foil
 - Au foil
 - Cu foil
 - Pt foil
- Spontaneous redox reaction takes place in:
 - Electrolytic cell
 - Galvanic cell
 - Voltaic cell
 - Both 'b' & 'c'
- The movement of ions in a solution, towards their respective electrodes is called:
 - Ionization
 - Electrolysis
 - Electrolytic conduction
 - All
- In Down's cell, cathode is made up of:
 - Graphite
 - Platinum
 - Iron
 - Pb
- When Zn electrode is connected to SHE, the following reaction takes place at SHE:
 - $H_{2(g)} \rightarrow 2H^+ + 2e^-$
 - $2H^+ + 2e^- \rightarrow H_{2(g)}$
 - $2H_2O + 2e^- \rightarrow H_{2(g)} + 2OH^-$
 - $2H_2O \rightarrow 2H_2 + O_2$
- The standard electrode potential is measured under the conditions:
 - 0.1 M solutions, $0^\circ C$, 1 atm pressure.
 - 0.1 M solutions, $25^\circ C$, 1 atm pressure.
 - 1 M solution, $0^\circ C$, 1 atm pressure.
 - 1 M solution, $25^\circ C$, 1 atm pressure.
- Which metal can displace Cu from $CuSO_4$ solution?
 - Fe
 - Ag
 - Hg
 - All

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- The oxidation number of oxygen in OF_2 is +2. Justify.
- Differentiate between metallic conduction and electrolytic conduction.
- Differentiate between electrolytic cell and Galvanic cell.
- What is Hall-Beroult process?
- What is SHE?
- What is emf of the cell? How is it calculated from electrochemical series?

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: What is Galvanic cell? Give composition and working of Galvanic cell.