

CHEMISTRY-11	Chapter#09-Second Half (9.5 – 9.8) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Melting point of ice can be lowered by the use of:
 (a) LiCl (b) BeCl₂ (c) NaCl (d) AgCl
- Quantitatively, solubility is the number of grams of _____ dissolved in 100 grams of the _____.
 (a) Solute, solvent (b) Solute, solution (c) Solvent, solution (d) None
- Solubility curve is the graphical representation between the solubility of a substance and:
 (a) Temperature (b) Pressure (c) Molality (d) Volume
- Which solution will cause greater elevation in boiling point of water?
 (a) 6 g urea in 1 kg water (b) 18 g glucose in 1 kg water
 (c) 34.2 g sucrose in 1 kg water (d) All equal
- The values of K_b and K_f depend upon:
 (a) The nature of solute (b) The nature of the solvent
 (c) The number of solute particles (d) Both 'a' & 'b'
- The presence of a solute in a liquid increases the _____ range of the solution.
 (a) Solid (b) Liquid (c) Gas (d) All
- On dissolving a salt in water, the temperature of the solution decreases and the container cools. What the salt may be?
 (a) KI (b) Li₂CO₃ (c) LiCl (d) Both 'b' & 'c'
- What is the correct order with respect to hydration energy of ions?
 (a) Cu²⁺ > Mg²⁺ > Ag⁺ (b) Mg²⁺ > Ag⁺ > Cu²⁺ (c) Ag⁺ > Mg²⁺ > Cu²⁺ (d) Mg²⁺ > Cu²⁺ > Ag⁺
- One mole of a solute requires _____ of water to reach the stage of infinite dilution.
 (a) 100-200 moles (b) 500-600 moles (c) 700-800 moles (d) 800-1000 moles
- One of the formulas given below is incorrect. Point out the incorrect one.
 (a) BaCl₂.2H₂O (b) Na₂CO₃.10H₂O (c) MgCl₂.6H₂O (d) AlCl₃.5H₂O

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(12)

- Define solubility curves. Name its two types.
- What are the applications of colligative properties?
- What is cryoscopic constant or molal freezing point constant?
- Why Beckmann thermometer is used to note the depression of the freezing point?
- Why hydration energy of Li⁺ ion is greater than Cs⁺ ion?
- Aqueous solution of CuSO₄ is acidic in nature. Justify.

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: Describe Landsburger's method for measuring the boiling point elevation.

Q.4: Define hydrolysis. Why the aqueous solution of NH₄Cl is acidic and that of CH₃COONa is basic?