

CHEMISTRY-11	Chapter#11-Second Half (11.4 – 11.6) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- In the method of large excess, the rate and order of the reaction is controlled by that reactant which is taken in:
 - Small amount
 - Excess amount
 - Both 'a' & 'b'
 - None
- If the rate equation of the reaction $2\text{NO} + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O} + \text{N}_2$ is $\text{Rate} = k[\text{H}_2]^1[\text{NO}]^2$. By tripling the concentration of NO, the rate becomes:
 - 3 times
 - 4 times
 - 8 times
 - 9 times
- The order of the reaction $2\text{NO} + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O} + \text{N}_2$ is:
 - 1
 - 3
 - 2
 - 0
- In daylight, the reaction between H_2 and Cl_2 is:
 - Slow
 - Moderate
 - Explosive
 - Negligible
- The relationship between temperature and the rate constant is called:
 - Rate law
 - Rate equation
 - Arrhenius equation
 - Raoult's law
- The Arrhenius constant "A", used in Arrhenius equation, depends upon:
 - The total number of collisions
 - The number of effective collisions
 - Both 'a' & 'b'
 - None
- The slope obtained by Arrhenius equation is calculated by:
 - Sin θ
 - Cosine θ
 - Tan θ
 - None
- HCl is oxidized to Cl_2 in the presence of:
 - Pt
 - CuCl_2
 - MnO_2
 - Ni
- Which one is the example of heterogeneous catalysis?
 - Oxidation of SO_2 in Lead chamber process
 - Acid hydrolysis of ester
 - Oxidation of NH_3 to form NO
 - None
- HCl and NH_3 combine to form white fumes of NH_4Cl in the presence of:
 - Pt
 - Cu
 - Ni
 - Moisture

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- Combustion occurs more rapidly in pure oxygen than in air, why?
- How does increase in temperature increases the rate of reaction?
- How does a catalyst alter the rate of a chemical reaction?
- A very small amount of catalyst may prove sufficient to carry out a chemical reaction. Explain with example.
- A catalyst is specific its action. Explain.
- What is meant by negative catalysis? Give an example.

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: What is half life period? Give example, also give its mathematical form.

Q.3: What is enzyme catalysis? Give an example. Also give any four characteristics of enzyme catalysis.