

CHEMISTRY-11	Chapter#10-Second Half (10.3-10.5) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Oxidation potential of SHE is:
(a) 0 (b) 1 (c) 2 (d) 4
- A single lead cell provides volts:
(a) 2 (b) 4 (c) 6 (d) 8
- When Zn electrode is connected to SHE, the following reaction takes place at Zn electrode.
(a) $Zn_{(s)} \rightarrow Zn^{2+} + 2e^-$ (b) $Zn^{2+} + 2e^- \rightarrow Zn_{(s)}$ (c) $Zn_{(s)} + 2e^- \rightarrow Zn^{2-}$ (d) $Zn^{2+} \rightarrow Zn^{4+}_{(s)} + 2e^-$
- Which electrode pair will produce maximum voltage?
(a) F_2/K (b) Cl_2/Zn (c) F_2/Li (d) I_2/K
- Silver battery is a type of _____ cell.
(a) Primary (b) Secondary (c) Tertiary (d) Electrolytic
- The electrolyte solution used in the lead accumulator is:
(a) 15% H_2SO_4 (b) 25% H_2SO_4 (c) 30% H_2SO_4 (d) 45% H_2SO_4
- In lead accumulator, salt bridge is not used because:
(a) Cathode and anode are made up of the same material.
(b) Electrolyte is same in both compartments.
(c) The same compound $PbSO_4$ is deposited on both cathode and anode.
(d) All
- The cathode in silver oxide battery is made up of:
(a) Zn (b) Cd (c) MnO_2 (d) Ag_2O
- The voltage of alkaline battery is:
(a) 1.2 V (b) 1.5 V (c) 1.4 V (d) 2 V
- Fuels cells convert _____ of fuel bond energy into electricity.
(a) 50% (b) 70% (c) 75% (d) 80%

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- Differentiate between single electrode potential and standard electrode potential.
- What is SHE?
- SHE acts as cathode when connected with Zn electrode while as anode when connected with Cu electrode. Why?
- Write electrode reactions occurring in lead storage battery.
- What are the electrode reactions in silver oxide battery?
- What are fuel cells? How do they generate electrical energy?

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: What is electrochemical series? Give its any four applications.

Q.4: Write a note on Nickel-Cadmium cell.