

CHEMISTRY-11	Chapter#10-Second Half (10.3-10.5) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Galvanic cells which can't be re-charged are called:
 (a) Diffused cells (b) Tertiary cells (c) Secondary cells (d) Primary cells
- In silver oxide battery, the cathode is made up of:
 (a) AgO (b) Ag₂O (c) Ag₂O₂ (d) Ag
- When Zn electrode is connected to SHE, the following reaction takes place at SHE:
 (a) $H_{2(g)} \rightarrow 2H^+ + 2e^-$ (b) $2H^+ + 2e^- \rightarrow H_{2(g)}$
 (c) $2H_2O + 2e^- \rightarrow H_{2(g)} + 2OH^-$ (d) $2H_2O \rightarrow 2H_2 + O_2$
- Which element has the highest value of standard reduction potential?
 (a) Li (b) H₂ (c) K (d) F₂
- The greater is the value of reduction potential, the greater is the reactivity of the:
 (a) Metal (b) Non-metal (c) Both 'a' & 'b' (d) None
- Dry cell is a type of _____ cell.
 (a) Primary (b) Secondary (c) Tertiary (d) Electrolytic
- The density of H₂SO₄ solution used in lead accumulator is:
 (a) 1 gcm⁻³ (b) 1.25 gcm⁻³ (c) 1.5 gcm⁻³ (d) 1.15 gcm⁻³
- The total cell voltage produced by the lead accumulator is:
 (a) 2 Voltas (b) 6 volts (c) 10 volts (d) 12 volts
- The anode in Nickle-Cadmium cell is made up of:
 (a) NiO₂ (b) Cd (c) MnO₂ (d) Ag₂O
- In fuel cells, water is produced at:
 (a) Cathode (b) Anode (c) Both 'a' & 'b' (d) None

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(12)

- What happens when a metal electrode is dipped in the solution of its own ions?
- What is emf of the cell? How is it calculated from electrochemical series?
- The standard oxidation potential of Zn is 0.76 V and its reduction potential is -0.76 V. Justify.
- Lead accumulator is a chargeable battery. Comment on it.
- What are the electrode reactions in silver oxide battery?
- Write down the reactions occurring in fuel cells.

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: Describe the construction and working of standard hydrogen electrode.

Q.3: Write a note on alkaline battery. Give its applications.