

CHEMISTRY-11	Chapter#08-Second Half(8.3 - 8.9) Test-2A		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The solubility product of AgCl is $2.0 \times 10^{-10} \text{ mole}^2 \text{ dm}^{-6}$. The maximum concentration of Ag^+ ions in the solution is:
 (a) $2.0 \times 10^{-10} \text{ mol dm}^{-3}$ (b) $1.41 \times 10^{-5} \text{ mol dm}^{-3}$ (c) $1.0 \times 10^{-10} \text{ mol dm}^{-3}$ (d) $4.0 \times 10^{-20} \text{ mol dm}^{-3}$
- The nature of milk is:
 (a) Acidic (b) Basic (c) Neutral (d) Normal
- The pH of human blood is maintained at:
 (a) 7.30 (b) 7.35 (c) 7.00 (d) 8.00
- $\text{pH} = \text{pK}_a + \log \frac{[\text{Salt}]}{[\text{Acid}]}$ is known as:
 (a) Einstein equation (b) Gibb's equation
 (c) Plank equation (d) Henderson equation
- If $\text{K}_a < 10^{-3}$, the acid is:
 (a) Strong (b) Moderately strong
 (c) Weak (d) Moderately weak
- The greater is the pK_a value, the _____ is the acid.
 (a) Weaker (b) Stronger (c) Moderate (d) None
- When KCl is added to the saturated solution of KClO_3 then:
 (a) The solubility of KCl decreases (b) The solubility of KClO_3 decreases
 (c) The solubility of KClO_3 increases (d) The solubility of KCl increases
- The pH of a buffer containing $\text{H}_2\text{CO}_3 + \text{Na}_2\text{CO}_3$ will be:
 (a) = 7 (b) >7 (c) <7 (d) None
- If K_{sp} is greater than ionic product of salt, the result is:
 (a) Un-saturation (b) Saturation (c) Precipitation (d) None

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(12)

- Calculate the pH of $10^{-3} \text{ mole dm}^{-3} \text{ HCl}$.
- What is Ostwald's dilution law?
- How can NaCl be purified through common ion effect?
- Why do we need buffers?
- How does the mixture of CH_3COOH and CH_3COONa give the acidic buffer?
- Calculate solubility product expression for Ag_2CrO_4 .

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: What is ionic product of water? How does its value vary with change in temperature?