

CHEMISTRY-11	Chapter#11-Second Half (11.4 – 11.6) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Arrhenius equation is represented as:
 - $A = k^{-E_a/RT}$
 - $A = k^{RT/-E_a}$
 - $k = Ae^{RT/-E_a}$
 - $k = Ae^{-E_a/RT}$
- The catalyst used for the reaction $HCOOH \rightarrow H_2 + CO_2$ is:
 - Copper
 - Alumina
 - Silica
 - Iron
- The thermal decomposition of N_2O is a _____ order reaction.
 - First
 - Second
 - Third
 - Zero
- The rate of reaction generally doubles with every _____ increase in temperature.
 - 10 K
 - 30 K
 - 20 K
 - 40 K
- The slope obtained by Arrhenius equation is equal to:
 - $\frac{E_a}{2.303 R}$
 - $-\frac{E_a}{2.303 R}$
 - $\frac{E_a}{2.303 RT}$
 - $-\frac{E_a}{2.303 RT}$
- The decomposition of $KClO_3$ occurs rapidly in the presence of:
 - Pt
 - $CuCl_2$
 - MnO_2
 - Ni
- 2.5 dm^3 of H_2 and 1.25 dm^3 of O_2 can be converted into water in the presence of only _____ of fine platinum powder.
 - 1 g
 - 0.1 g
 - 1 mg
 - 0.1 g
- Tetraethyl lead, used to reform petroleum, acts as a:
 - Poison
 - Negative catalyst
 - Auto catalyst
 - Promoter
- The role of enzyme as catalyst resembles:
 - Inorganic homogeneous catalysts:
 - Organic homogeneous catalyst
 - Inorganic heterogeneous catalyst
 - Organic heterogeneous catalyst

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- Define half-life period. How is it used to determine the order of reaction?
- How does surface area affect the rate of the reaction? Give an example.
- What is Arrhenius equation? What is its importance? OR How is temperature related with the rate constant?
- Write two characteristics of a catalyst.
- What is a promoter or activator? Give an example.
- Enzymes are highly specific in action. Justify.

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: What is catalysis? Differentiate between homogeneous and heterogeneous catalysis with one example in each.