

CHEMISTRY-11	Chapter#10-Second Half (10.3-10.5) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- If a strip of Cu metal is placed in a solution of FeSO_4 :
 - Cu will be deposited
 - Fe is precipitated out
 - Cu and Fe both dissolve
 - No reaction takes place
- The best reducing agent is:
 - F^{-1}
 - Cl^{-1}
 - Br^{-1}
 - I^{-1}
- The voltage of Nickel Cadmium Cell:
 - 1 V
 - 1.2 V
 - 1.4 V
 - 1.6 V
- When connected to Zn electrode, SHE acts as:
 - Cathode
 - Anode
 - Both
 - None
- The standard electrode potential is measured under the conditions:
 - 0.1 M solutions, 0°C , 1 atm pressure.
 - 0.1 M solutions, 25°C , 1 atm pressure.
 - 1 M solution, 0°C , 1 atm pressure.
 - 1 M solution, 25°C , 1 atm pressure.
- The greater is the value of standard reduction potential, the greater is its tendency to act as:
 - An oxidizing agent
 - A reducing agent
 - Both as oxidizing and reducing agent
 - None of the above
- Which metal can release H_2 from steam?
 - Fe
 - Cu
 - Ag
 - Au
- Which metal can displace Zn from ZnSO_4 solution?
 - Cu
 - Fe
 - Pb
 - Mg
- The anode in alkaline battery is made up of:
 - Zn
 - Cd
 - MnO_2
 - Ag_2O

Maximum Marks: 16 **(SUBJECTIVE TYPE)** Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- SHE acts as cathode when connected with Zn electrode while as anode when connected with Cu electrode. Why?
- What is electrochemical series?
- Fe can displace Cu from CuSO_4 but Zn does not displace Mg from MgSO_4 solution. Why?
- Differentiate between primary and secondary cell.
- What is NICAD? Write electrode reactions occurring in it.
- What are the advantages of fuel cells?

SECTION-II

NOTE: Attempt All Questions: (04)

Q.3: Write a comprehensive note on lead accumulator with its discharging and recharging process.