

CHEMISTRY-11	Chapter#08-Second Half(8.3 - 8.9) Test-1B		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The pH of 10^{-3} mole dm^{-3} of an aqueous solution of H_2SO_4 is:
 (a) 3.0 (b) 2.7 (c) 2.0 (d) 1.5
- The value of K_w at 25°C is:
 (a) 0.11×10^{-14} (b) 0.30×10^{-14} (c) 1×10^{-14} (d) 3×10^{-14}
- The pH of the gastric juice is:
 (a) 2.0 (b) 3.0 (c) 3.5 (d) 5.6
- Acid having $K_a > 1$ will be:
 (a) Weak (b) Very weak (c) Moderate (d) Stronger
- By adding NH_4Cl to NH_4OH solution, the ionization of NH_4OH :
 (a) Increases (b) Decreases
 (c) Remains same (d) Increases 100 times
- The solubility of KClO_3 in water is suppressed by adding:
 (a) KClO_3 (b) NaCl (c) KMnO_4 (d) KCl
- $\text{pH} + \text{pOH} = ?$
 (a) K_w (b) $\text{p}K_a$ (c) $\text{p}K_w$ (d) $\text{p}K_b$
- Common ion effect is the application of:
 (a) Law of mass action (b) Le-Chatelier's Principle
 (c) Buffer (d) All
- If K_{sp} is smaller than the ionic product of salt, the result is:
 (a) Un-saturation (b) Saturation (c) Precipitation (d) None

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(12)

- Define pH and pOH. How are they related with $\text{p}K_w$?
- What do you mean by percentage ionization of acid?
- Define $\text{p}K_a$ and $\text{p}K_b$.
- Define common ion effect?
- Define buffer capacity.
- What is solubility product? Derive solubility product expression for PbCl_2 .

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: Calculate the pH of buffer solution which 0.11 molar CH_3COONa and 0.09 molar acetic acid solutions are present. K_c for CH_3COOH is 1.85×10^{-5} .