

CHEMISTRY-11	Chapter#10-Second Half (10.3-10.5) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Stronger the oxidizing agent, greater is the:
 - Oxidation potential
 - Reduction potential
 - Redox potential
 - E.M.F. of cell
- A typical car battery of 12 volts has number of cells:
 - 4
 - 3
 - 6
 - 2
- Percentage of H_2SO_4 solution used in lead accumulator is:
 - 40%
 - 25%
 - 30%
 - 50%
- Electrode potential is actually the potential difference, set up between:
 - Anode and cathode.
 - Anode compartment and cathode compartment.
 - The electrode and the solution of its ions.
 - Atoms and the electrons on the electrode.
- The oxidation potential of standard hydrogen electrode (SHE) is taken as:
 - 0
 - +1
 - 1
 - +10
- When Cu electrode is connected to SHE, the following reaction takes place at Cu electrode.
 - $Cu_{(s)} \rightarrow Cu^{2+} + 2e^-$
 - $Cu^{2+} + 2e^- \rightarrow Cu_{(s)}$
 - $Cu_{(s)} + 2e^- \rightarrow Cu^{2-}$
 - $Cu^{2+} \rightarrow Cu^{4+}_{(s)} + 2e^-$
- Which element has the lowest value of standard reduction potential?
 - Li
 - H_2
 - K
 - F_2
- Which metals will react with acids to liberate H_2 gas?
 - Cu, Ag
 - Au, Pt
 - Ca, Mg
 - Both 'a' & 'c'
- The cells which can't be recharged are called:
 - Primary cells
 - Secondary cells
 - Electrolytic cells
 - Electrochemical cell

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(12)

- Differentiate between single electrode potential and standard electrode potential.
- What is SHE?
- How does electrochemical series predict about the feasibility of a chemical reaction?
- Na and K can displace hydrogen from acids but Pt, Pd and Cu cannot. Why?
- A porous plate and salt bridge is not required in lead accumulator. Why?
- Write electrode reactions occurring in alkaline battery (dry cell).

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: Write a note on fuel cells.