

CHEMISTRY-11	Chapter#10-First Half (10.1-10.2) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- In H_2O_2 , the oxidation state of oxygen is:
(a) +1 (b) -1 (c) +2 (d) -2
- Oxidation number of Cl in $Ca(ClO_3)_2$ is:
(a) -1 (b) +5 (c) +3 (d) +1
- Which one is not an electrolyte?
(a) Aqueous NaCl (b) Cu metal (c) Molten NaCl (d) H_2SO_4
- The cell concerned with the conversion of electrical energy into chemical energy is called:
(a) Electrolytic cell (b) Galvanic cell (c) Voltaic cell (d) Both 'b' & 'c'
- In super oxides, the oxidation number of O is:
(a) 0 (b) $-\frac{1}{2}$ (c) -1 (d) -2
- Consider the equation: $K_2Cr_2O_7 + HCl \rightarrow KCl + CrCl_3 + Cl_2$. What is the correct statement with reference to the above equation?
(a) $K_2Cr_2O_7$ is reduced to $CrCl_3$ (b) $K_2Cr_2O_7$ is oxidized to $CrCl_3$
(c) HCl is reduced to Cl_2 (d) HCl is oxidized to $CrCl_3$
- The movement of ions in a solution, towards their respective electrodes is called:
(a) Ionization (b) Electrolysis
(c) Electrolytic conduction (d) All
- During the electrolysis of aqueous solution of $AgNO_3$ and HNO_3 , using silver electrodes, silver dissolves at:
(a) Cathode (b) Anode (c) Both 'a' & 'b' (d) None
- During the purification process of copper in electrolytic cell, pure copper is obtained at:
(a) Cathode (b) Anode (c) Both (d) None
- In Galvanic cell, the electrons flow externally from _____ electrode to _____ electrode.
(a) Zn, Cu (b) Zn, Zn (c) Cu, Zn (d) Cu, Cu

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- Define oxidation number. What is oxidation number of elements in free state?
- Calculate the oxidation number of Mn in $KMnO_4$ and MnO_2 .
- Differentiate between oxidation and reduction.
- Differentiate between electrolytic cell and Galvanic cell.
- How is anodized aluminum prepared?
- A salt bridge maintains the electrical neutrality in the cell. Explain.

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: State rules for assigning oxidation number of elements with examples.

Q.4: Write four industrial applications of electrolysis.