

CHEMISTRY-11	Chapter#10-First Half (10.1-10.2) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Which of the following statements is not correct about galvanic cell?
 - Anode is negatively charged
 - Reduction occurs at anode
 - Cathode is positively charged
 - Reduction occurs at cathode
- In Na_2O_2 , the oxidation state of oxygen is:
 - 2
 - +2
 - 1
 - +1
- Which process is used for the extraction of Al?
 - Castner-Kellner process
 - Hall-Beroult process
 - Thermite process
 - Combustion process
- The cell concerned with the conversion of chemical energy into electrical energy is called:
 - Electrolytic cell
 - Galvanic cell
 - Voltaic cell
 - Both 'b' & 'c'
- Which element can never show positive oxidation state?
 - N
 - O
 - F
 - Both 'b' & 'c'
- In balancing of redox equations by ion electron method in basic medium, O is balanced by:
 - H^+
 - OH^-
 - H_2O
 - None
- Reduction occurs at:
 - Anode
 - Cathode
 - Both 'a' & 'b'
 - None
- Salt bridge in Galvanic cell contains an aqueous solution of _____ in gel.
 - NaCl
 - KCl
 - NaNO_3
 - KNO_3
- The voltage produced in Galvanic cell is:
 - 1.1 V
 - 1.4 V
 - 1.2 V
 - 1.5 V
- The oxidation potential of Zn is:
 - +0.76 V
 - +0.34 V
 - 0.76 V
 - 0.34 V

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

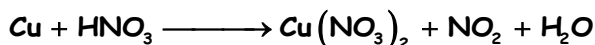
Q.2: Give brief answers to the following questions: (12)

- The oxidation number of oxygen in OF_2 is +2. Justify.
- Calculate the oxidation number of Cr in K_2CrO_4 and CrCl_3 .
- Define electrochemical cells.
- How is caustic soda obtained in Nelson cell?
- What is Hall-Beroult process?
- A salt bridge maintains the electrical neutrality in the cell. Explain.

SECTION-II

NOTE: Attempt All Questions: (08)

Q.3: Balance the following equation by oxidation number method:



Q.4: Describe the electrolysis of molten sodium chloride and a concentrated solution of sodium chloride.