

CHEMISTRY-11	Chapter#09-First Half (9.0 - 9.4) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Molarity of pure water is:
 - 1
 - 18
 - 55.5
 - 6
- An azeotropic mixture of two liquids boils at a lower temperature than either of them when:
 - It is saturated.
 - It shows positive deviation from Raoult's law.
 - It shows negative deviation from Raoult's law.
 - It is metastable.
- Relative lowering of vapour pressure is equal to:
 - Mole fraction of solute
 - Mole fraction of solvent
 - Molarity
 - Molality
- 2 g of NaCl dissolved in 20 g of water makes _____ solution.
 - 2% w/w
 - 9% w/w
 - 5% w/w
 - 10% w/w
- The molarity of water is maximum at:
 - 0°C
 - 25°C
 - 4°C
 - 100°C
- Cheese is an example of:
 - Solid into solid solution
 - Solid into liquid solution
 - Liquid into solid solution
 - Liquid into liquid solution
- The critical solution temperature of water-aniline system is:
 - 65.9°C
 - 49.1°C
 - 167.0°C
 - 176.0°C
- The vapour pressure of the solvent above the solution is directly proportional to the:
 - Mole fraction of solute
 - Mole fraction of solvent
 - Both 'a' & 'b'
 - None
- If the enthalpy of solution of a mixture of two liquids is negative, it will show _____ deviation from Raoult's law.
 - Positive
 - Negative
 - Both 'a' & 'b'
 - None

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- Define phase.
- One molal aqueous solution of urea is more dilute than one molar solution. But the number of particles in them are same. Explain?
- The sum of mole fractions of all the components is always equal to unity. Explain.
- What are conjugate solutions?
- Relative lowering of vapour pressure is independent of temperature. Explain.
- Differentiate between zeotropic and azeotropic mixtures.

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: Define non-ideal solutions and explain positive deviation with the help of a graph.