

CHEMISTRY-11	Chapter#11-First Half (11.0 – 11.3) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- If the rate equation of a reaction $2A + B \longrightarrow$ Products is, $\text{rate} = k [A]^2[B]$, and A is present in large excess, then order of reaction is:
 - 1
 - 2
 - 3
 - None
- The unit of the rate constant is the same as that of the rate of reaction in:
 - First order reaction
 - Second order reaction
 - Zero order reaction
 - Third order reaction
- The rate of the reaction $A \rightarrow B$, can be expressed as:
 - $\frac{d[A]}{dt}$
 - $\frac{-d[B]}{dt}$
 - $\frac{-d[A]}{dt}$
 - $\frac{-d[A]}{-dt}$
- The units of rate constant for third order reaction are:
 - Moles $\text{dm}^{-3}\text{s}^{-1}$
 - Moles $^{-1}\text{dm}^3\text{s}^{-1}$
 - Moles $^{-2}\text{dm}^6\text{s}^{-1}$
 - s^{-1}
- The half-life period for the disintegration of radioactive ${}_{92}^{225}\text{U}$ is:
 - 7.1×10^8 thousand years
 - 7.1×10^8 million years
 - 7.1×10^8 billion years
 - 7.1×10^8 years
- For a zero-order reaction, if the concentration of the reactants is doubled, the half-life:
 - Also becomes double
 - Becomes half
 - Becomes four times
 - Remains same
- The reaction $\text{NO}_2 + \text{CO} \rightarrow \text{NO} + \text{CO}_2$ is _____ order with respect to CO.
 - First
 - Second
 - Third
 - Zero
- The necessary condition for a collision to be effective:
 - Activation energy
 - Proper orientation
 - Both 'a' & 'b'
 - None
- For exothermic reaction, the E_a of reverse reaction is _____ that of forward reaction.
 - Greater than
 - Smaller than
 - Equal to
 - All

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(12)

- Differentiate between average rate and instantaneous rate.
- Photochemical reactions are usually zero order reactions. Explain.
- The slowest step is the rate determining step. Explain.
- How rate of reaction is determined by electrical conductivity method?
- What are the conditions for a collision to be effective?
- Define activation energy.

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: Define order of reaction and give one example of first, second and third order reactions.