CHEMISTRY-11	Chapter#10-First Half (10.1-10.2) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1. **The cathode reaction in the electrolysis of dil. H_2SO_4 with Pt electrodes is:**
   (a) Reduction (b) Oxidation
   (c) Both oxidation and reduction (d) Neither oxidation or reduction
2. **Oxidation number of chromium in $Cr_2O_7^{-2}$ is:**
   (a) +3 (b) +4 (c) +5 (d) +6
3. **The oxidation number of Cl in $HClO_4$ is:**
   (a) +2 (b) +3 (c) +5 (d) +7
4. **Non-spontaneous redox reaction takes place in:**
   (a) Electrolytic cell (b) Galvanic cell (c) Voltaic cell (d) Both 'b' & 'c'
5. **Which oxidation number can't be shown by oxygen?**
   (a) -1 (b) -2 (c) +2 (d) None of these
6. **In balancing of redox equations by ion electron method in basic medium, H is balanced by:**
   (a) H^+ (b) OH^- (c) H_2O (d) None
7. **During the electrolysis of aqueous $CuSO_4$ solution, _____ is deposited at anode.**
   (a) Cu (b) SO_2 (c) H_2 (d) O_2
8. **During electrolytic purification of copper, anode is made up of:**
   (a) Pure copper (b) Impure copper (c) Graphite (d) Platinum
9. **In Galvanic cell, anode is _____ charged.**
   (a) Positively (b) Negatively (c) Neutrally (d) Both 'a' & 'b'

Maximum Marks: 16 **(SUBJECTIVE TYPE)** Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

i. Define electrochemistry.
ii. Calculate the oxidation number of underlined element. a) $H_3\underline{P}O_3$ b) $Ca(\underline{Cl}O_3)_2$.
iii. Differentiate between oxidation and reduction.
iv. Explain the difference between ionization and electrolysis.
v. How is anodized aluminum prepared?
vi. What is a salt bridge? What are its functions?

SECTION-II

NOTE: Attempt All Questions: (04)

Q.3: Define electrolysis? Explain the electrolysis of very dilute solution of $NaNO_3$.