CHEMISTRY-11	Chapter#11-First Half (11.0 – 11.3) Test-1		
	Name:	Class:	ID:
Date: / / Time Allowed: 40 Min.	Marks Total: **25**	Marks Obtained:	

Maximum Marks: 09   **(OBJECTIVE TYPE)**   Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1. **In zero order reaction, the rate is independent of:**
   - (a) Temperature of reaction
   - (b) Concentration of reactants
   - (c) Concentration of products
   - (d) None of these

2. **Specific rate constant is equal to rate of reaction, when concentration of reactants is:**
   - (a) Zero
   - (b) Four
   - (c) Three
   - (d) Unity

3. **All radioactive disintegration nuclear reactions are of:**
   - (a) First order
   - (b) Second order
   - (c) Third order
   - (d) Zero order

4. **Near the end of the reaction, the instantaneous rate is _____ than the average rate.**
   - (a) Lower
   - (b) Higher
   - (c) Equal
   - (d) None

5. **The units of rate constant for second order reaction are:**
   - (a) $\text{Moles dm}^{-3}\text{s}^{-1}$
   - (b) $\text{Moles}^{-1}\text{dm}^3\text{s}^{-1}$
   - (c) $\text{Moles}^{-2}\text{dm}^6\text{s}^{-1}$
   - (d) s^{-1}

6. **The half-life period of the _____ order reaction is independent of the initial concentration of the reactants.**
   - (a) First
   - (b) Third
   - (c) Second
   - (d) Zero

7. **If the number of ions changes during a chemical reaction, the rate of such reaction can be determined by:**
   - (a) Dilatometric method
   - (b) Electrical conductivity
   - (c) Spectrometric method
   - (d) Optical rotation method

8. **Polarimeter is used in _____ method for measuring the rate of the reaction.**
   - (a) Spectrometric
   - (b) Refractometric
   - (c) Electrical conductivity
   - (d) Optical rotation

9. **The hydrolysis of ester in the presence of acid is _____ reaction.**
   - (a) First order
   - (b) Second order
   - (c) Zero order
   - (d) Pseudo first order

Maximum Marks: 16   **(SUBJECTIVE TYPE)**   Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)
   i. Define reaction kinetics.
   ii. Define rate equation or rate law.
   iii. Define order of reaction? Give an example.
   iv. The radioactive decay is always a first order reaction. Explain.
   v. What are reaction intermediates?
   vi. Write names of the physical methods to determine the rate of reaction.

SECTION-II

NOTE: Attempt All Questions: (04)

Q.3: Explain the concept of activation energy and activated complex.