

CHEMISTRY-11	Chapter#09(Complete) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 40	Marks Obtained:	
Time Allowed: 75 Min.			

Maximum Marks: 08

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- An aqueous solution of ethanol in water may have vapour pressure:
 - Equal to that of water
 - Equal to that of ethanol
 - More than that of water
 - Less than that of water
- Which one of the following pair of liquids is not completely miscible?
 - Alcohol and Ether
 - Benzene and Cyclohexane
 - Phenol and Water
 - Alcohol and water
- _____ salt gives neutral solution when dissolved in water:
 - NH_4Cl
 - AlCl_3
 - CuSO_4
 - Na_2SO_4
- Which one is affected by temperature?
 - Molarity
 - Molality
 - Mole fraction
 - All
- Diethyl ether dissolves water up to the extent:
 - 1.2%
 - 5%
 - 6.5%
 - 30%
- The solubility of NaCl at 0°C is:
 - 14.3 g
 - 37.5 g
 - 75.4 g
 - None
- The elevation in boiling point (ΔT_b) is directly related to:
 - Molarity
 - Molality
 - Both
 - None
- Which salt has exothermic heat of solution?
 - LiCl
 - KNO_3
 - NaCl
 - NH_4NO_3

Maximum Marks: 32

(SUBJECTIVE TYPE)

Time Allowed: 65 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(20)

- Differentiate between solvent and solute.
- The sum of mole fractions of all the components is always equal to unity. Explain.
- Explain the effect of temperature on phenol-water system.
- Relative lowering of vapour pressure is independent of temperature. Explain.
- What is positive deviation from Raoult's law?
- What is fractional crystallization?
- Why 6 g urea, 18 g glucose and 34.2 g of sucrose produce same lowering in vapour pressure?
- What are the conditions to observe the colligative properties?
- What is the use of ethylene glycol in an automobile radiator?
- Lattice energy of ionic solids is always higher than molecular solids. Why?

SECTION-II

NOTE: Attempt All Questions:

(12)

Q.3: Define the following terms: i) Mole fraction ii) Molarity iii) Parts per million iv) Hydration.

Q.4: What is solubility? Explain solubility curves.

Q.5: What is hydration and hydrolysis? Explain with two examples.