

CHEMISTRY-11	Chapter#09(Complete) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 60 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- A solution of glucose is 10% w/v. The volume in which 1 g mole of it is dissolved will be:
 (a) 1 dm³ (b) 1.8 dm³ (c) 200 cm³ (d) 900 cm³
- In azeotropic mixture showing positive deviation from Raoult's law, the volume of the mixture is:
 (a) Slightly more than the total volume of the components.
 (b) Slightly less than the total volume of the components.
 (c) Equal to the total volume of the components.
 (d) None of these
- The mass of glucose required to prepare 1 dm³ of 20% glucose solution is:
 (a) 18 g (b) 180 g (c) 36 g (d) 200 g
- The critical solution temperature of methanol-cyclohexane system is:
 (a) 65.9°C (b) 49.1°C (c) 167.0°C (d) 176.0°C
- The azeotropic mixture of water and HCl contains HCl:
 (a) 95.5% (b) 20.24% (c) 79.76% (d) 4.5%
- CuSO₄ will give _____ solution in water.
 (a) Acidic (b) Basic (c) Neutral (d) None

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 50 Min.

SECTION-I

- Q.2: Give brief answers to the following questions: (16)**
- One molal aqueous solution of urea is more dilute than one molar solution. But the number of particles in them are same. Explain?
 - Define critical solution temperature or the upper consolute temperature.
 - Differentiate between zeotropic and azeotropic mixtures.
 - What is discontinuous solubility curve? Give one example.
 - What is ebullioscopic constant or molal boiling point constant?
 - NaCl and KNO₃ are used to lower the melting point of ice. Explain.
 - Why hydration energy of Li⁺ ion is greater than Cs⁺ ion?
 - What is water of crystallization? Give two examples.

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: Explain Raoult's law when both components in the solution are volatile.

Q.4: Describe Landsburger's method for measuring the boiling point elevation.