

CHEMISTRY-11	Chapter#09(Complete) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 60 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- 18 g glucose is dissolved in 90 g water. The relative lowering of Vap. pressure is equal to:
 - 1/5
 - 5.1
 - 1/51
 - 6
- Colligative properties are the properties of:
 - Dilute solutions which behave as nearly ideal solutions.
 - Concentrated solutions which behave as nearly non-ideal solutions.
 - Both 'a' & 'b'
 - Neither 'a' nor 'b'
- The sum of mole fractions of gases in a mixture of gases is:
 - Always more than 1
 - May be less or more than 1
 - Always less than 1
 - Always 1
- An aqueous solution of methanol in water has vapor pressure:
 - Equal to that of water
 - Less than that of water
 - Equal to that of methanol
 - More than that of water
- The amount of NaOH required to prepare 250 cm³ of 1 M solution in grams is:
 - 10
 - 15
 - 20
 - 25
- Which one is the example of partially miscible liquids?
 - Alcohol + Water
 - Alcohol + Ether
 - Water + Ether
 - Water + CS₂

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 50 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(16)

- Differentiate between dilute solution and concentrated solution.
- Define molality. Give its mathematical expression.
- What are conjugate solutions?
- What is Raoult's law? Give its three definitions with mathematical expressions.
- Define solubility curves. Name its two types.
- What are the applications of colligative properties?
- Why hydration energy of Na⁺ ion is smaller than Li⁺ ion?
- Why the aqueous solution of NH₄Cl is acidic?

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: Differentiate between ideal and non-ideal solutions.

Q.4: How is depression in freezing point of a solution is measured by Beckmann's method?