

CHEMISTRY-11	Chapter#09(Complete) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 60 Min.			

Maximum Marks: 06 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Molarity of pure water is:
 (a) 1 (b) 18 (c) 55.5 (d) 6
- The molal boiling point constant is the ratio of the elevation in boiling point to:
 (a) Molarity (b) Molality
 (c) Mole fraction of solvent (d) Mole fraction of solute
- Upper consolute temperature for water-phenol system is:
 (a) 150°C (b) 65.9°C (c) 120°C (d) 130°C
- A thermometer used in Landsberger's method can read up to:
 (a) 0.1 K (b) 0.01°F (c) 0.01 K (d) 0.01°C
- Water of crystallization of oxalic acid is:
 (a) 2 (b) 10 (c) 7 (d) 6
- Which one of the following gives acidic solution when dissolved in H₂O?
 (a) NaCl (b) Na₂SO₄ (c) NH₄Cl (d) CH₃COONH₄

Maximum Marks: 24 **(SUBJECTIVE TYPE)** Time Allowed: 50 Min.

SECTION-I

- Q.2: Give brief answers to the following questions: (16)**
- Define a solution.
 - Define molarity. Give its mathematical expression.
 - What is solubility principle?
 - Differentiate between ideal and non-ideal solutions.
 - How is the solubility of a substance determined?
 - What are colligative properties? Why are they called so? Name some of them.
 - Define hydration energy of ions.
 - Aqueous solution of Na₂CO₃ is alkaline in nature, why?

SECTION-II

NOTE: Attempt All Questions: (08)

- Q.3: What is Raoult's Law? Give its three statements.**
- Q.4: What is molal boiling point constant? How can we determine the molar mass of an unknown solid by elevation in boiling point method?**