

CHEMISTRY-11	Chapter#11(Complete) Test-6		
	Name:	Class:	ID:
Date: / /	Marks Total: 40	Marks Obtained:	
Time Allowed: 75 Min.			

Maximum Marks: 08

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The reaction which occurs at a moderate rate:
 - Between AgNO_3 and NaCl solutions.
 - Fermentation of sugars.
 - Hydrolysis of ester.
 - Rusting of iron.
- The average rate becomes equal to the instantaneous rate when the time interval approaches to:
 - Zero
 - Maximum
 - Minimum
 - Infinity
- The overall order of reaction $\text{NO} + \text{O}_3 \rightarrow \text{NO}_2 + \text{O}_2$ is:
 - 1
 - 2
 - 0
 - Pseudo first
- The half-life period for the decomposition of N_2O_5 at 45°C is 24 minutes. If we start with 1 mole dm^{-3} of N_2O_5 , then after 96 minutes how much of its mass will be left?
 - 0.0625 mole
 - 0.0750 mole
 - 0.0250 mole
 - 0.0125 mole
- The reaction intermediate formed during the reaction $\text{NO}_2 + \text{CO} \rightarrow \text{NO} + \text{CO}_2$ is:
 - N_2O_4
 - N_2O
 - N_2O_5
 - NO_3
- Dilatometric method is used for measuring the rate of reactions, involving:
 - Temperature change
 - Conductivity changes
 - Pressure changes
 - Volume changes
- In darkness, the reaction between H_2 and Cl_2 is:
 - Slow
 - Moderate
 - Explosive
 - Negligible
- The oxidation of NH_3 to NO is catalyzed by:
 - Ni
 - Pd
 - Pt
 - All

Maximum Marks: 32

(SUBJECTIVE TYPE)

Time Allowed: 65 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (20)

- Differentiate between average rate and instantaneous rate.
- Rate of reaction changes every moment but rate constant k remains the same, why?
- Photochemical reactions are usually zero order reactions. Explain.
- What are reaction intermediates?
- Describe method of large excess for finding the order of reaction?
- How does increase in temperature increases the rate of reaction?
- Differentiate between homogenous and heterogeneous catalyses.

- viii. A catalyst may be changed physically at the end of reaction. Explain with example.
- ix. What are the types of catalytic poisoning?
- x. What is optimum pH of enzyme?

SECTION-II

NOTE: Attempt All Questions:

(12)

Q.3: What is half life period? Give example, also give its mathematical form.

Q.4: How light and surface area affect the rate of chemical reaction.

Q.5: How does Arrhenius equation help us to calculate the activation energy of a reaction.