

CHEMISTRY-11	Chapter#11(Complete) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 60 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- In zero order reaction, the rate is independent of:
 - Temperature of reaction
 - Concentration of reactants
 - Concentration of products
 - None of these
- All radioactive disintegration nuclear reactions are of:
 - First order
 - Second order
 - Third order
 - Zero order
- Glucose can be converted into ethanol by an enzyme:
 - Lipase
 - Zymase
 - Sucrase
 - Urease
- The overall order of the reaction $2\text{FeCl}_3 + \text{KI} \rightarrow 2\text{FeI}_2 + 6\text{KCl} + \text{I}_2$ is:
 - 1
 - 2
 - 3
 - 0
- Polarimeter is used in _____ method for measuring the rate of the reaction.
 - Spectrometric
 - Refractometric
 - Electrical conductivity
 - Optical rotation
- Arrhenius equation explains the effect of _____ on the rate constant of a reaction.
 - Temperature
 - Concentration
 - Pressure
 - Activation energy

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 50 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (16)

- Define reaction kinetics.
- The radioactive decay is always a first order reaction. Explain.
- What are reaction intermediates?
- What is refractometric method for measuring the rate of reaction?
- Why powdered Al reacts more rapidly with cold NaOH than Al foil with warm NaOH?
- How is temperature related with the rate constant?
- What is poisoning of catalyst? Explain with examples.
- Give two characteristics of enzyme catalysis.

SECTION-II

NOTE: Attempt All Questions:

(08)

- Q.3: What is half life period? Explain with examples. Also write formulas to calculate the half life periods of first, second and third order reactions.**
- Q.4: Write four characteristics of a catalyst.**