CHEMISTRY-11	Chapter#11(Complete) Test-2		
	Name:	Class:	ID:
Date: / / Time Allowed: 60 Min.	Marks Total: **30**	Marks Obtained:	

Maximum Marks: 06   **(OBJECTIVE TYPE)**   Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1. If the rate equation of a reaction $2A + B \longrightarrow$ Products is, rate $= k[A]^2[B]$, and A is present is large excess, then order of reaction is:
    (a) 1   (b) 2   (c) 3   (d) None

2. The catalyst used for the reaction $HCOOH \rightarrow H_2 + CO_2$ is:
    (a) Copper   (b) Alumina   (c) Silica   (d) Iron

3. If a reaction occurs in several steps, the _____ step will be the rate determining step.
    (a) First   (b) Last   (c) Fastest   (d) Slowest

4. The units of rate constant for zero order reaction are:
    (a) Moles $dm^{-3}s^{-1}$ (b) Moles$^{-1}dm^3s^{-1}$ (c) Moles$^{-2}dm^6s^{-1}$ (d) s^{-1}

5. Half-life period of any reaction is inversely proportional to the initial concentration of the reactant raised to power _____ than the order of the reaction.
    (a) One more   (b) One less   (c) Two more   (d) Two less

6. The compounds of _____ behave as poisons to many metallic catalysts.
    (a) P and As   (b) S and Si   (c) P and Al   (d) S and As

Maximum Marks: 24   **(SUBJECTIVE TYPE)**   Time Allowed: 50 Min.

SECTION-I

Q.2: Give brief answers to the following questions: **(16)**

  i. Define rate of reaction. Give its units?

 ii. Differentiate between rate and rate constant of a reaction.

iii. What are pseudo first order reactions? Explain with an example.

 iv. What are the conditions for a collision to be effective?

  v. Name the factors which affect the rate of reaction.

 vi. Define catalyst and catalysis.

vii. A catalyst is specific its action. Explain.

viii. Define autocatalysis with one example.

SECTION-II

NOTE: Attempt All Questions: **(08)**

Q.3: Explain the concept of activation energy and activated complex.

Q.4: What is enzyme catalysis? Give an example. Also give any four characteristics of enzyme catalysis.