

CHEMISTRY-11	Chapter#11(Complete) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 60 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The unit of the rate constant is the same as that of the rate of reaction in:
 - First order reaction
 - Second order reaction
 - Zero order reaction
 - Third order reaction
- The order of decomposition of nitrogen pentaoxide ($2\text{N}_2\text{O}_5 \rightarrow 2\text{N}_2\text{O}_4 + \text{O}_2$) is:
 - First order
 - Second order
 - Third order
 - Zero order
- Half life period of ${}_{92}\text{U}^{235}$ is:
 - 710 million years
 - 720 million years
 - 810 million years
 - 820 million years
- Arrhenius equation is represented as:
 - $A = k^{-E_a/RT}$
 - $A = k^{RT/-E_a}$
 - $k = Ae^{RT/-E_a}$
 - $k = Ae^{-E_a/RT}$
- The hydrolysis of sugar is catalyzed by:
 - Urease
 - Invertase
 - Zymase
 - Diastase
- The necessary condition for a collision to be effective:
 - Activation energy
 - Proper orientation
 - Both 'a' & 'b'
 - None

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 50 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(16)

- Define reaction kinetics.
- Define rate equation or rate law.
- Write names of the physical methods to determine the rate of reaction.
- Define half-life period. How is it used to determine the order of reaction?
- How does light affect the rate of photochemical reactions?
- What is poisoning of catalyst? Explain with examples.
- What is meant by negative catalysis? Give an example.
- Give two characteristics of enzyme catalysis.

SECTION-II

NOTE: Attempt All Questions:

(08)

- Q.3: Define order of reaction and give one example of first, second and third order reactions.
- Q.4: What is catalysis? Differentiate between homogeneous and heterogeneous catalysis with one example in each.