CHEMISTRY-11	Chapter#10 (Complete) Test-4		
	Name:	Class:	ID:
Date: / /	**Marks Total: 40**	**Marks Obtained:**	
Time Allowed: 75 Min.			

Maximum Marks: 08 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1. **If the salt bridge is not used between two half cells, then the voltage:**
   (a) Decreases rapidly (b) Decreases slowly (c) Does not change (d) Drops to zero
2. **Oxidation No. of C in $C_{12}H_{22}O_{11}$ is:**
   (a) Zero (b) -6 (c) 16 (d) 12
3. **When one metal is deposited on the surface of the other by the process of electrical current, it is called:**
   (a) Electrolysis (b) Electroplating
   (c) Electrolytic refining (d) Electrolytic
4. **The voltage of Nickel Cadmium Cell:**
   (a) 1 V (b) 1.2 V (c) 1.4 V (d) 1.6 V
5. **In balancing of redox equations by ion electron method in basic medium, O is balanced by:**
   (a) H^+ (b) OH^- (c) H_2O (d) None
6. **During the electrolysis of aqueous $CuSO_4$ solution, _____ is deposited at anode.**
   (a) Cu (b) SO_2 (c) H_2 (d) O_2
7. **Salt bridge in Galvanic cell contains an aqueous solution of _____ in gel.**
   (a) NaCl (b) KCl (c) $NaNO_3$ (d) KNO_3
8. **The greater is the value of reduction potential, the greater is the reactivity of the:**
   (a) Metal (b) Non-metal (c) Both 'a' & 'b' (d) None

Maximum Marks: 32 **(SUBJECTIVE TYPE)** Time Allowed: 65 Min.

SECTION-I

Q.2: Give brief answers to the following questions: **(20)**

i. Calculate the oxidation number of Cr in K_2CrO_4 and $CrCl_3$.
ii. Differentiate between oxidation and reduction.
iii. Differentiate between electrolytic cell and Galvanic cell.
iv. What is a salt bridge? What are its functions?
v. SHE acts as cathode when connected with Zn electrode while as anode when connected with Cu electrode. Why?
vi. How does electrochemical series predict about the feasibility of a chemical reaction?
vii. Na and K can displace hydrogen from acids but Pt, Pd and Cu cannot. Why?
viii. The standard oxidation potential of Zn is 0.76 V and its reduction potential is -0.76 V. Justify.
ix. A porous plate and salt bridge is not required in lead accumulator. Why?
x. What are the electrode reactions in silver oxide battery?

SECTION-II

NOTE: Attempt All Questions: **(12)**

Q.3: Write four industrial applications of electrolysis.
Q.3: Describe the construction and working of standard hydrogen electrode.
Q.3: Write a note on fuel cells.