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| CHEMISTRY-11 | Chapter#10 (Complete) Test-3 | | |
| | Name: | Class: | ID: |
| Date: / / | Marks Total: 30 | Marks Obtained: | |
| Time Allowed: 60 Min. | | | |

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Stronger the oxidizing agent, greater is the:
 - Oxidation potential
 - Reduction potential
 - Redox potential
 - E.M.F. of cell
- The oxidation number of O in OF₂ is:
 - 2
 - +2
 - 1
 - +1
- Cathode in NICAD cell is:
 - Ag₂O
 - NiO₂
 - Cd
 - Zn
- Non-spontaneous redox reaction takes place in:
 - Electrolytic cell
 - Galvanic cell
 - Voltaic cell
 - Both 'b' & 'c'
- Select the one which is oxidation half reaction.
 - Cl⁻ → Cl₂
 - MnO₄⁻ → Mn²⁺
 - MnO₄⁻ → MnO₂
 - NO₃¹⁻ → 2NO₂
- Fuels cells convert _____ of fuel bond energy into electricity.
 - 50%
 - 70%
 - 75%
 - 80%

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 50 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(16)

- The oxidation number of oxygen in OF₂ is +2. Justify.
- Calculate the oxidation number of Sulphur in SO₄²⁻.
- Differentiate between electrolytic cell and Galvanic cell.
- How is anodized aluminum prepared?
- Differentiate between single electrode potential and standard electrode potential.
- What is emf of the cell? How is it calculated from electrochemical series?
- Differentiate between primary and secondary cell.
- Write down the reactions occurring in fuel cells.

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: Define electrolysis? Explain the electrolysis of very dilute solution of NaNO₃.

Q.4: Write a comprehensive note on lead accumulator with its discharging and recharging process.