

<b>CHEMISTRY-11</b>	<b>Chapter#10 (Complete) Test-2</b>		
	Name:	Class:	ID:
Date: / /	<b>Marks Total: 30</b>	<b>Marks Obtained:</b>	
Time Allowed: 60 Min.			

Maximum Marks: 06

**(OBJECTIVE TYPE)**

Time Allowed: 10 Min.

**NOTE:** Tick The Correct Option:

- Which of the following statements is not correct about galvanic cell?
  - Anode is negatively charged
  - Reduction occurs at anode
  - Cathode is positively charged
  - Reduction occurs at cathode
- In  $\text{Na}_2\text{O}_2$ , the oxidation state of oxygen is:
  - 2
  - +2
  - 1
  - +1
- In balancing of redox equations by ion electron method in basic medium, H is balanced by:
  - $\text{H}^+$
  - $\text{OH}^-$
  - $\text{H}_2\text{O}$
  - None
- The movement of ions in a solution, towards their respective electrodes is called:
  - Ionization
  - Electrolysis
  - Electrolytic conduction
  - All
- Which metal can release  $\text{H}_2$  from steam?
  - Fe
  - Cu
  - Ag
  - Au
- Which one is used in portable computers and cordless razors?
  - Alkaline battery
  - Silver oxide battery
  - NICAD
  - Fuel cells

Maximum Marks: 24

**(SUBJECTIVE TYPE)**

Time Allowed: 50 Min.

**SECTION-I**

Q.2: Give brief answers to the following questions:

(16)

- Calculate the oxidation number of Mn in  $\text{KMnO}_4$  and  $\text{MnO}_2$ .
- Differentiate between metallic conduction and electrolytic conduction.
- Define electrochemical cells.
- What is Hall-Beroult process?
- A salt bridge maintains the electrical neutrality in the cell. Explain.
- What is emf of the cell? How is it calculated from electrochemical series?
- Differentiate between primary and secondary cell.
- Write electrode reactions occurring in alkaline battery (dry cell).

**SECTION-II**

**NOTE:** Attempt All Questions:

(08)

Q.3: Balance the following equation by oxidation number method:



Q.4: What is standard hydrogen electrode (SHE)? How is it used to measure the electrode potential of copper?