CHEMISTRY-11	Chapter#07(Complete-Smart Syllabus) Test-2		
	Name:	Class:	ID:
Date: / /	**Marks Total: 25**	**Marks Obtained:**	
Time Allowed: 40 Min.			

Maximum Marks: 09 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1) In endothermic reaction, the heat content of the:
(a) Products is more than that of reactants (b) Reactants is more than that of products
(c) Both 'a' & 'b' (d) Reactants and products are equal

2) For a given process, the heat changes at constant pressure (q_p) and at constant volume (q_v) are related to each other as:
(a) $q_p = q_v$ (b) $q_p < q_v$ (c) $q_p > q_v$ (d) $q_p = \frac{q_v}{2}$

3) For the reaction: $H^+ + OH^- \longrightarrow H_2O$, the change in enthalpy is called heat of:
(a) Reaction (b) Combustion (c) Solution (d) Neutralization

4) We are studying the reaction of Zn with $CuSO_4$ to form $ZnSO_4$. The flask in which the reaction is taking place is:
(a) System (b) Surroundings (c) Boundary (d) None

5) A di-atomic molecule shows _____ motion.
(a) Only translational (b) Translational and vibrational
(c) Translational, vibrational and rotational (d) None

6) 25°C is the standard temperature in:
(a) Thermochemistry (b) Thermodynamics (c) Electrochemistry (d) All

7) Which salt will raise the temperature of the solution when dissolved in water?
(a) Na_2CO_3 (b) NaCl (c) NH_4Cl (d) All

8) The specific heat of water is:
(a) 4.2 JK^{-1} mol^{-1} (b) 4.2 J K^{-1} g^{-1} (c) 4.2 J K^{-1} (d) 4.2 $kJK^{-1}g^{-1}$

9) Which one is the enthalpy of atomization of bromine?
(a) 108 kJ mol^{-1} (b) 112 kJ mol^{-1} (c) 121 kJ mol^{-1} (d) 218 kJ mol^{-1}

Maximum Marks: 16 **(SUBJECTIVE TYPE)** Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: **(12)**

i. What is state and state function? Explain with example.
ii. What is meant by pressure volume work?
iii. Why is it necessary to mention the physical states of reactants and products in a thermochemical reaction?
iv. What do you mean by enthalpy of atomization?
v. Why can the enthalpy of formation of CCl_4 not be measured directly?
vi. State Hess's law of constant heat summation.

SECTION-II

NOTE: Attempt All Questions: **(04)**

Q.3: State first law of thermodynamics? How does it explain that $\Delta H = q_p$.