CHEMISTRY-11	Chapter#06 (Complete-Smart Syllabus) Test-3		
	Name:	Class:	ID:
Date: / /	**Marks Total: 30**	**Marks Obtained:**	
Time Allowed: 50 Min.			

Maximum Marks: 10 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1) An ionic compound $A^+ B^-$ is most likely to be formed when:

(a) The ionization energy of A is high and electron affinity of B is low.
(b) The ionization energy of A is low and electron affinity of B is high.
(c) Both the ionization energy of A and electron affinity of B are high.
(d) Both the ionization energy of A and electron affinity of B are low.

2) The value of third ionization energy of Mg is:

(a) 1450 $kJmol^{-1}$ (b) 7730 $kJmol^{-1}$ (c) 7850 $kJmol^{-1}$ (d) 1890 $kJmol^{-1}$

3) CsF has ionic character:

(a) 60% (b) 80% (c) 92% (d) 100%

4) The bond angle between two H-S bonds in H_2S is:

(a) 180° (b) 104.5° (c) 109.5° (d) 92°

5) Bond order of O_2^{2-} is:

(a) Zero (b) One (c) Two (d) Three

6) Which one is the correct order of the radii of H-atoms?

(a) Covalent > atomic > anionic > cationic (b) Atomic > anionic > covalent > cationic
(c) Anionic > covalent > atomic > cationic (d) Anionic > atomic > covalent > anionic

7) The least electronegative element in the periodic table is:

(a) H (b) Li (c) Cs (d) I

8) A regular tetrahedron has _____ corners.

(a) 4 (b) 6 (c) 8 (d) 2

9) The mode of hybridization in $BeCl_2$ is:

(a) sp^3 (b) sp^2 (c) sp (d) None

10) Which molecular orbital will remain empty in O_2 molecule?

(a) $\sigma 2p_x$ (b) $\sigma^* 2p_x$ (c) $\pi^* 2p_y$ (d) $\pi^* 2p_z$

Maximum Marks: 20 **(SUBJECTIVE TYPE)** Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions: **(12)**

i. Bond distance is the compromise distance between two atoms. Explain.
ii. No bond in chemistry is 100% ionic. Explain.
iii. The bond angle in NF_3 is smaller than in NH_3. Justify.
iv. Pi bonds are more diffused than sigma bonds. Explain.
v. Differentiate between bonding and anti-bonding molecular orbitals.
vi. Why is O_2 paramagnetic?

SECTION-II

NOTE: Attempt All Questions: **(08)**

Q.3: Define coordinate covalent bond. Explain with examples.

Q.4: What is hybridization? Explain the geometry of ethene on the basis of sp^2 hybrization.