

CHEMISTRY-11	Chapter#06 (Complete-Smart Syllabus) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The number of bonds in nitrogen molecule is:
 (a) One σ and one π (b) One σ and two π (c) Three σ only (d) Two σ and one π
- The octet rule is not followed by:
 (a) NH_3 (b) Cl_2 (c) CCl_4 (d) PCl_5
- The shape of SnCl_2 molecule is:
 (a) Linear (b) Angular (c) Trigonal planar (d) Tetrahedral
- In sp^2 hybridization, the hybrid orbitals are oriented at angle of:
 (a) 107.5° (b) 109.5° (c) 120° (d) 180°
- Ionization energy is inversely proportional to:
 (a) Atomic radius (b) Shielding effect (c) Nuclear charge (d) Both 'a' & 'b'
- Ionic character of NaCl :
 (a) 72% (b) 92% (c) 100% (d) 50%
- VSEPR theory was developed by:
 (a) Lewis and Kossel (b) Nyholm and Gillespie
 (c) Sidgwick and Powell (d) Pauling
- Molecular orbital theory considers the whole _____ as a single unit.
 (a) Atom (b) Molecule (c) Radical (d) Both 'a' & 'b'
- The bond order for He_2 molecule is:
 (a) Zero (b) 1 (c) 2 (d) -1

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

- Q.2: Give brief answers to the following questions:** (12)
- Define chemical bond. Give the cause of chemical bond formation.
 - Why cationic radius is smaller than the parent atom?
 - Why is second I.E. greater than the first one?
 - Define co-ordinate covalent bond with the help of two examples.
 - Why sigma bond is stronger than a pi bond?
 - What is bond order? How is bond order calculated?

SECTION-II

NOTE: Attempt All Questions: (04)

- Q.3: Calculate the bond order of O_2 molecule by making energy level diagram. Also show that it is paramagnetic.**