

CHEMISTRY-11	Chapter#05 (Complete-Smart Syllabus) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Bohr model of atom is contradicted by:
 - Plank's quantum theory
 - Dual nature of matter
 - Heisenberg's uncertainty principal
 - All of the above
- Orbitals having same energy are called:
 - Hybrid orbitals
 - Valence orbitals
 - Degenerate orbitals
 - d-orbitals
- Cathode rays can be generated at the pressure of:
 - 1 torr
 - 0.1 torr
 - 0.01 torr
 - 0.001 torr
- The limiting line of Balmer series lies in the region:
 - Visible
 - Ultraviolet
 - Near I.R.
 - Far I.R.
- At which voltage, the gas discharge tube glows under reduced pressure?
 - 1000-5000 V
 - 5000-10000 V
 - 10000-15000 V
 - All of the above
- Which sub-atomic particles are used in the treatment of cancer?
 - Protons
 - Electrons
 - Neutrons
 - α -particles
- When electron is promoted to higher orbit, its velocity:
 - Increases
 - Decreases
 - Remains same
 - None
- The spectral lines of Balmer series arise when the electrons in H-atoms jump from higher orbits to:
 - 1st orbit
 - 2nd orbit
 - 3rd orbit
 - 4th orbit
- Heisenberg's uncertainty principle can be best explained with the help of:
 - Compton's effect
 - Photoelectric effect
 - Zeeman effect
 - Stark effect
- The number of electrons in d-subshell is:
 - 2
 - 6
 - 10
 - 14

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- Cathode rays have reducing effect. Explain.
- How were neutrons discovered?
- Justify that the distance gaps between orbits go on increasing from lower to higher orbits.
- Define spectrum. Give its two types.
- How was dual nature of electron verified?
- Why are d-orbitals called five fold degenerate orbitals?

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: Give different postulates of Bohr's atomic model.

Q.4: What are quantum numbers? Explain azimuthal quantum number and magnetic quantum number.

