

CHEMISTRY-11	Chapter#05 (Complete-Smart Syllabus) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- When electric current is passed through hydrogen gas in a discharge tube, it glows:
 - Red
 - Green
 - Yellow
 - Bluish
- The nature of positive rays depends on:
 - The nature of electrodes
 - The nature of the discharge tube
 - The nature of the residual gas
 - All of the above
- Cathode rays strike alumina and produce a _____ color.
 - Red
 - Blue
 - Yellow
 - Green
- Fast neutrons move with an energy:
 - 1 eV
 - 1.4 eV
 - 1.2 Mev
 - 0.2 Mev
- Which rays are non-material?
 - Cathode rays
 - Positive rays
 - X-rays
 - All
- Nucleus was discovered by:
 - J. J. Thomson
 - Rutherford
 - Millikan
 - Bohr
- The velocity of electron will be maximum in:
 - E_1
 - E_2
 - E_3
 - E_∞
- The volume of space in which there is _____ chance of finding an electron is called atomic orbital.
 - 90%
 - 95%
 - 99%
 - 100%
- Quantum number values for 2p orbital are:
 - $n = 2, \ell = 1$
 - $n = 1, \ell = 2$
 - $n = 1, \ell = 0$
 - $n = 2, \ell = 0$

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- How was it inferred that cathode rays are material particles?
- Give reason for the production of positive rays?
- Calculate the mass of an electron when its e/m is $1.758 \times 10^{11} \text{ C kg}^{-1}$.
- What are the defects in Bohr's model?
- What is Zeeman effect?
- What are degenerate orbitals?

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: Describe Millikan's oil droplet method for the determination of charge of electron.