

CHEMISTRY-11	Chapter#04 (Complete-Smart Syllabus) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- London dispersion forces are the only forces present among the:
 - Molecules of water in liquid state.
 - Atoms of helium in gaseous state at high temperature.
 - Molecules of solid iodine.
 - Molecules of hydrogen chloride gas.
- Dipole-dipole forces are _____ effective as covalent bond.
 - 1%
 - 5%
 - 20%
 - 50%
- What is the correct order for boiling points?
 - $\text{HF} > \text{H}_2\text{O} > \text{NH}_3 > \text{HCl}$
 - $\text{H}_2\text{O} > \text{HF} > \text{HCl} > \text{NH}_3$
 - $\text{H}_2\text{O} > \text{HF} > \text{NH}_3 > \text{HCl}$
 - $\text{H}_2\text{O} > \text{NH}_3 > \text{HF} > \text{HCl}$
- The small regions of orderly arrangement in amorphous solids are called:
 - Unit cell
 - Lattice sites
 - Crystallites
 - Crystallographic elements
- MgO is isomorphic to:
 - NaF
 - NaNO_3
 - ZnSO_4
 - K_2CrO_4
- The unit cell angle present between a and b is called:
 - α
 - β
 - γ
 - Φ

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (16)

- What are dipole-dipole forces?
- Define polarizability.
- H_2O is a liquid while H_2S is a gas at room temperature. Explain.
- State biological application of liquid crystals.
- Why is graphite anisotropic in electrical conductivity?
- Define transition temperature with an example.
- Transition temperature is lower than melting point, why?
- Define crystal lattice.

SECTION-II

NOTE: Attempt All Questions:

(08)

- Q.3: Define and explain London forces. Describe the factors affecting the London dispersion forces.**
- Q.4: Differentiate between isomorphism and polymorphism with suitable examples.**