

<b>CHEMISTRY-11</b>	<b>Chapter#04 (Complete-Smart Syllabus) Test-1</b>		
	Name:	Class:	ID:
Date: / /	<b>Marks Total: 25</b>	<b>Marks Obtained:</b>	
Time Allowed: 40 Min.			

Maximum Marks: 09

**(OBJECTIVE TYPE)**

Time Allowed: 10 Min.

**NOTE:** Tick The Correct Option:

- Acetone and chloroform are soluble into each other due to:
  - Intermolecular hydrogen bonding
  - Ion-dipole interaction
  - Instantaneous dipole
  - All of the above
- The molecules of  $CO_2$  in dry ice form the:
  - Ionic crystals
  - Covalent crystals
  - Molecular crystals
  - Any type of crystals
- Dipole-dipole forces are present among:
  - Molecules of Iodine
  - Atoms of Neon in gaseous state
  - Chloroform molecules
  - $CCl_4$  molecules
- Density of  $H_2O$  is maximum at:
  - $0^\circ C$
  - $2^\circ C$
  - $-1^\circ C$
  - $4^\circ C$
- Cholesteryl benzoate turns into milky liquid at:
  - $144^\circ C$
  - $145^\circ C$
  - $146^\circ C$
  - $147^\circ C$
- Which one has dipole-dipole forces?
  - $H_2O$
  - HCl
  - $C_6H_6$
  - All
- Down the group, polarizability:
  - Increases
  - Decreases
  - Remains same
  - All are possible
- The number of hydrogen bonds per HF molecule is:
  - 1
  - 2
  - 3
  - 4
- Graphite is a good conductor \_\_\_\_\_ to the layers.
  - Parallel
  - Across
  - Perpendicular
  - Both 'a' & 'b'

Maximum Marks: 16

**(SUBJECTIVE TYPE)**

Time Allowed: 30 Min.

**SECTION-I**

- Q.2: Give brief answers to the following questions:** (12)
- Differentiate between intermolecular and intramolecular forces.
  - What is the role of London dispersion forces in the liquefaction of noble gases?
  - Why is HF weaker than all the halogen acids?
  - Give four uses of liquid crystals.
  - What is habit of crystal? Can habit be changed?
  - Differentiate between polymorphism and allotropy.

**SECTION-II**

**NOTE:** Attempt All Questions:

(04)

- Q.3: What is hydrogen bonding? Discuss hydrogen bonding in biological compounds.**