

CHEMISTRY-11	Chapter#03 (Complete) Test-2 (SMART SYLLABUS)		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Which of the followings, will have the same number of molecules at STP?
 - 280 cm³ of CO₂ and 280 cm³ of N₂O
 - 11.2 dm³ of O₂ and 32 g of O₂
 - 44 g of CO₂ and 11.2 dm³ of CO
 - 28 g of N₂ and 5.6 dm³ of oxygen
- Absolute zero is equal to:
 - 273°C
 - 273°C
 - 0°C
 - 273 K
- The graphical representation of Charles' law is called:
 - An isotherm
 - An isobar
 - A negative slope
 - A hyperbolic curve
- If the pressure on the gas is reduced to one half and the temperature is doubled, the volume will become:
 - Half
 - Double
 - Four times
 - Same
- The partial pressure of O₂ in inhaled air is:
 - 760 torr
 - 593 torr
 - 159 torr
 - 116 torr
- The Sun is a _____ ball of plasma.
 - 1 million km
 - 1.5 million km
 - 2 million km
 - 2.5 million km

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: GIVE BRIEF ANSWERS TO THE FOLLOWING QUESTIONS: (16)

- What happens to the positions of isotherms, when they are plotted at higher temperatures?
- Justify that the volume of given mass of a gas becomes theoretically zero at -273°C.
- Describe different scales of thermometry.
- Calculate the value of R in SI units.
- Deep sea divers or scuba divers do not use normal air in breathing, why?
- Define mean square velocity and root mean square velocity.
- Derive Charles' law from KMT.
- What are the characteristics of plasma?

SECTION-II

NOTE: Attempt All Questions:

(08)

State Boyle's law. Give its experimental verification.

State and explain Dalton's law of partial pressures. Derive its mathematical expression for calculating partial pressure of a gas.