

CHEMISTRY-11	Chapter#03 (Complete) Test-1 (SMART SYLLABUS)		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- If absolute temperature of a gas is doubled and the pressure is reduced to one half, the volume of the gas will be:
 - Remain unchanged
 - Increase four times
 - Reduce to $\frac{1}{4}$
 - Be doubled
- The value of general gas constant "R" in $\text{dm}^3 \text{ atm K}^{-1} \text{ mol}^{-1}$ is:
 - 62.4
 - 0.0821
 - 62400
 - 1.987
- The temperature of natural plasma is about:
 - 20,000°C
 - 10,000°C
 - 5000°C
 - 1000°C
- The graph between P and 1/V, at constant temperature, is:
 - A straight line
 - An isotherm
 - A negative slope
 - A parabolic curve
- The formula for density of a gas at a given temperature and pressure is:
 - $d = \frac{PV}{RT}$
 - $d = \frac{Pm}{RT}$
 - $d = \frac{PM}{RT}$
 - $d = \frac{RT}{PM}$
- Which one is an example of artificial plasma?
 - Auroras
 - Neon signs
 - Lightning bolts
 - All

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: GIVE BRIEF ANSWERS TO THE FOLLOWING QUESTIONS: (16)

- What do you mean by the term isotherm?
- Define absolute Zero.
- Convert: (a) -40°F into Centigrade scale. (b) 40°C into Fahrenheit scale.
- State Avogadro's law.
- Do you think that 1 mole of H_2 and 1 mole of NH_3 at 0°C and 1 atm pressure will have Avogadro's number of particles?
- At higher altitude, pilots feel uncomfortable breathing, why?
- Derive Boyle's law from KMT.
- Give important uses of plasma.

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: Prove general gas equation ($PV=nRT$).

Q.4: Give eight postulates of KMT (Kinetic Molecular Theory).