

CHEMISTRY-11	Chapter#04-Seond Half (4.4-4.8) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Diamond is a bad conductor because:
 - It has a tight structure.
 - It has a high density.
 - There are no free electrons present in the crystal of diamond to conduct electricity.
 - Is transparent to light.
- Allotropy is the property of:
 - Compound
 - Element
 - Atom
 - Mixture
- Crystal of diamond is:
 - Ionic
 - Covalent
 - Molecular
 - Metallic
- Crystalline solids have _____ melting points.
 - Sharp
 - High
 - Low
 - Wide
- Which ion has tetrahedral shape?
 - CrO_4^{2-}
 - $\text{Cr}_2\text{O}_7^{2-}$
 - CO_3^{2-}
 - NO_3^-
- Ionic solids are:
 - Malleable
 - Ductile
 - Brittle
 - All
- The number of formula units present per unit cell of NaCl is:
 - 1
 - 4
 - 6
 - 8
- C-C bond length in diamond is:
 - 1.54 Å
 - 1.34 Å
 - 1.20 Å
 - 1.397 Å
- The lattice points in metallic solid are:
 - Atoms
 - Molecules
 - Anions
 - Cations
- The empty spaces between the metal atoms are called:
 - Voids
 - Crevice
 - Interstices
 - All

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(12)

- What are crystallites?
- The crystals showing isomorphism mostly have same atomic ratio. Explain.
- What are unit cell dimensions or crystallographic elements?
- Why ionic crystalline solids have high melting points?
- Why NaCl and CsCl have different structures?
- Metals are malleable and ductile. Why?

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: Differentiate between isomorphism and polymorphism with suitable examples.

Q.4: Explain the structure of diamond.