

CHEMISTRY-11	Chapter#03 - Second Half (3.7 to 3.11) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- If 'a' and 'b' are zero for certain gas, then gas is:
 - Ideal
 - Non-ideal
 - Real
 - May be any diatomic gas
- Kinetic equation was derived by:
 - Bernoulli
 - Maxwell
 - Clausius
 - Boltzmann
- Every gas consists of small particles called:
 - Atoms
 - Ions
 - Molecules
 - All
- In solids, temperature is the measure of _____ kinetic energy of the molecules.
 - Translational
 - Rotational
 - Vibrational
 - All
- In Linde's method, air is compressed to about:
 - 100 atm
 - 200 atm
 - 300 atm
 - 400 atm
- Ideal gases obey gas laws at:
 - Low temperature and low pressure.
 - Low temperature and high pressure.
 - High temperature and low pressure.
 - All temperatures and pressures.
- Which gas shows least deviation from ideal behavior?
 - H₂
 - He
 - N₂
 - CO₂
- Plasma was discovered by:
 - Van der Waals
 - J.J. Thomson
 - Clausius
 - William Crooks
- The Sun is a _____ ball of plasma.
 - 1 million km
 - 1.5 million km
 - 2 million km
 - 2.5 million km
- Which one is an example of natural plasma?
 - Auroras
 - Fluorescent bulbs
 - Neon signs
 - None

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

- Q.2: Give brief answers to the following questions: (12)**
- Derive Boyle's law from KMT.
 - Define critical temperature, critical pressure, and critical volume.
 - Polar gases have higher critical temperature than non-polar gases. Why?
 - H₂ and He are ideal at room temperature but SO₂ and Cl₂ are non-ideal Explain.
 - Why was the pressure correction done by van der Waals'?
 - What is natural and artificial plasma? OR What are the types of plasma?

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.0: Give eight postulates of KMT (Kinetic Molecular Theory).

Q.0: Explain non-ideal behaviour of gases.