CHEMISTRY-11	Chapter#04-Seond Half (4.4-4.8) Test-3		
	Name:	Class:	ID:
Date: / /	**Marks Total: 30**	**Marks Obtained:**	
Time Allowed: 50 Min.			

Maximum Marks: 10 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1. **Amorphous solids:**
   (a) Have sharp melting points.
   (b) Undergo clean cleavage when cut with knife.
   (c) Have perfect arrangement of atoms.
   (d) Can possess small regions of orderly arrangement of atoms.
2. **Which of the followings is a pseudo solid?**
   (a) CaF_2 (b) Glass (c) NaCl (d) All
3. **Transition temperature of KNO_3 is:**
   (a) 13.2°C (b) 95.5°C (c) 128°C (d) 32.2°C
4. **Cleavage is an _____ property.**
   (a) Isotropic (b) Anisotropic (c) Chemical (d) Optical
5. **Isomorphs have different _____ properties.**
   (a) Physical (b) Chemical (c) Both 'a' & 'b' (d) None
6. **The shape of SO_4^{2-} ion is:**
   (a) Tetrahedral (b) Octahedral (c) Square planar (d) Pyramidal
7. **Select the one which has triclinic crystals.**
   (a) $K_2Cr_2O_7$ (b) NH_4Br (c) $BaSO_4$ (d) $Na_2SO_4.10H_2O$
8. **Which one is not the property of ionic solids?**
   (a) High stability (b) High volatility (c) Hardness (d) High m.p. & b.p.
9. **Atomic solid is another name of:**
   (a) Ionic solids (b) Molecular solids (c) Covalent solids (d) Metallic solids
10. **Pauling explained the metallic bond as:**
    (a) Ionic bond (b) Delocalized covalent bond
    (c) Localized covalent bond (d) Co-ordinate covalent bond

Maximum Marks: 20 **(SUBJECTIVE TYPE)** Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions: **(12)**

i. Why is graphite anisotropic in electrical conductivity?
ii. What is habit of crystal? Can habit be changed?
iii. Transition temperature is lower than melting point, why?
iv. Ionic solids do not conduct electricity in solid state but conduct electricity in molten or solution form. Why?
v. Iodine dissolves readily in tetrachloromethane. Why?
vi. Freshly cut metals give lustrous surface. Explain.

SECTION-II

NOTE: **Attempt All Questions:** **(04)**

Q.3: Define the following with examples:

i) Amorphous solids ii) Habit of crystals iii) Allotropy iv) Transition temperature

Q.4: Write four properties of metallic crystals.