

CHEMISTRY-11	Chapter#07-Second Half (7.4.1 – 7.5.1) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The net heat change in a chemical reaction is same, whether it is brought about in two or more different ways in one or several steps. It is known as:
 - Henry's law
 - Joule's principle
 - Hess's law
 - Law of conservation of energy
- In endothermic reaction, ΔH is taken as:
 - Positive
 - Zero
 - Negative
 - May be any value
- For the reaction: $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$, the change in enthalpy is called:
 - Heat of reaction
 - Heat of neutralization
 - Heat of formation
 - Heat of combustion
- In a Bomb Calorimeter, the reactions are carried out at constant:
 - Pressure
 - Temperature
 - Volume
 - Enthalpy
- Which one is always exothermic?
 - ΔH°_f
 - ΔH°_{at}
 - ΔH°_c
 - ΔH°_{sol}
- When one mole of H_2 reacts with half mole $\text{O}_{2(g)}$ to form one mole gaseous water $\text{H}_2\text{O}_{(g)}$, the enthalpy change is:
 - 242.2 kJ
 - 285.8 kJ
 - 484 kJ
 - +285.8 kJ
- The heat required to raise the temperature of one gram of a substance through 1°C or 1 K is called:
 - Specific heat
 - Heat capacity
 - Molar specific heat
 - All
- Born-Haber cycle helps us to calculate the _____ of binary ionic compounds.
 - Bond energies
 - Hydration energies
 - Lattice energies
 - Formation enthalpies
- Enthalpy of combustion of graphite is $-393.51 \text{ kJ mol}^{-1}$ while that of diamond is $-395.41 \text{ kJ mol}^{-1}$. Can you guess which one is more stable?
 - Diamond
 - Graphite
 - Both 'a' & 'b'
 - None

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- What is a thermochemical equation? Give three examples. What information does it convey?
- Why is it necessary to mention the physical states of reactants and products in a thermochemical reaction?
- What do you mean by enthalpy of atomization?
- Define standard enthalpy of solution. Give examples.
- How do we determine ΔH in the laboratory for food and fuel etc.
- Differentiate between specific heat and heat capacity.

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: State and explain Hess's law of constant heat summation with an examples.