

CHEMISTRY-11	Chapter#04-Seond Half (4.4-4.8) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The molecules of CO_2 in dry ice form the:
 - Ionic crystals
 - Covalent crystals
 - Molecular crystals
 - Any type of crystals
- The highest value of lattice energy is for which one of these ionic compounds?
 - NaI
 - NaF
 - NaBr
 - NaCl
- Iodine is:
 - An ionic solid
 - A covalent solid
 - A metallic solid
 - A molecular solid
- If two different compounds exist in same crystalline form, the phenomenon is called:
 - Isotropy
 - Isomorphism
 - Polymorphism
 - Allotropy
- The allotropic forms of tin (Sn) are cubic and _____.
 - Tetragonal
 - Hexagonal
 - Rhombic
 - Trigonal
- The hexagonal crystal system has the unit cell dimensions:
 - $a = b = c; \alpha \neq \beta \neq \gamma \neq 90^\circ$
 - $a = b \neq c; \alpha = \beta = 90^\circ, \gamma = 120^\circ$
 - $a \neq b \neq c; \alpha = \beta = \gamma = 90^\circ$
 - $a \neq b \neq c; \alpha = \gamma = 90^\circ, \beta \neq 90^\circ$
- Ionic solids do not conduct electricity in:
 - Solid state
 - Molten state
 - Solution form
 - All
- The number of Cl^- ions per unit cell of NaCl is:
 - 1
 - 2
 - 4
 - 6
- Select the one which is a molecular solid:
 - NaCl
 - Sugar
 - Cu
 - AlN

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- Differentiate between crystalline and amorphous solids.
- Define anisotropy with example.
- Define unit cell. Give one example.
- Ionic crystals are highly brittle. Justify.
- Diamond is hard and electrical insulator. Why?
- In the closest packing of atoms of metals, only 74% space is occupied. How?

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: What are molecular solids? Give their properties and examples.