

<b>CHEMISTRY-11</b>	<b>Chapter#03 - Second Half (3.7 to 3.11) Test-2</b>		
	Name:	Class:	ID:
Date: / /	<b>Marks Total: 25</b>	<b>Marks Obtained:</b>	
Time Allowed: 40 Min.			

Maximum Marks: 09

**(OBJECTIVE TYPE)**

Time Allowed: 10 Min.

**NOTE:** Tick The Correct Option:

- A real gas obeying van der Waals equation will resemble ideal gas if:
  - Both "a" and "b" are large
  - Both "a" and "b" are small
  - "a" is small and "b" is large
  - "a" is large and "b" is small
- The temperature of natural plasma is about:
  - 20,000°C
  - 10,000°C
  - 5000°C
  - 1000°C
- Which is the correct expression for kinetic equation?
  - $PV = \frac{1}{3} Mn\overline{C^2}$
  - $PV = \frac{1}{3} mNc^2$
  - $PV = \frac{1}{3} MNc^2$
  - $PV = \frac{1}{3} mn\overline{C^2}$
- In liquids and gases, temperature is the measure of average \_\_\_\_\_ kinetic energy of the molecules.
  - Translational
  - Vibrational
  - Rotational
  - All
- Critical temperature is the \_\_\_\_\_ temperature at which a substance can exist in liquid state.
  - Minimum
  - Maximum
  - Average
  - Absolute
- For an ideal gas, the value of compressibility factor is \_\_\_\_\_ under all conditions.
  - 1
  - 2
  - 3
  - 0
- The SI units of van der Waals constant 'a' are:
  - $Nm^{-4} mol^{-2}$
  - $Nm^{-2} mol^{-2}$
  - $Nm^{+4} mol^{-2}$
  - $Nm^{+2} mol^{-2}$
- Plasma contains:
  - Electrons
  - Positive ions
  - Neutral particles
  - All of above
- Radioactive contamination is thought to be controlled by the use of:
  - High energy plasma
  - Low energy plasma
  - Lasers
  - Metastable molecules

Maximum Marks: 16

**(SUBJECTIVE TYPE)**

Time Allowed: 30 Min.

**SECTION-I**

Q.2: Give brief answers to the following questions:

(12)

- Describe law of distribution of velocities.
- Derive Charles' law from KMT.
- Define compressibility factor.
- Water vapours don't behave ideally at 273 K.
- What is b? Why is it greater than the actual volume of the gas molecules?
- Give important uses of plasma.

**SECTION-II**

**NOTE:** Attempt All Questions:

(04)

Q.3: What is Joule-Thomson effect? Explain Linde's method for the liquefaction of gases.