CHEMISTRY-11	Chapter#05-Second Half (5.6-5.9) Test-1		
	Name:	Class:	ID:
Date: / /	**Marks Total: 25**	**Marks Obtained:**	
Time Allowed: 40 Min.			

Maximum Marks: 09 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1) In the ground state of an atom, the electron is present:

(a) In the nucleus (b) In the second shell

(c) Nearest to the nucleus (d) Farthest from the nucleus

2) de-Broglie equation is represented by:

(a) $h = \frac{\lambda}{mv}$ (b) $m = \frac{h}{\lambda v}$ (c) $m = \frac{\lambda}{hv}$ (d) $\lambda = \frac{h}{mv}$

3) When 5d orbital is complete, then entering electron goes into:

(a) 6s (b) 6p (c) 6d (d) 6f

4) X-rays were discovered by:

(a) Roentgen (b) Heisenberg (c) Moseley (d) Schrodinger

5) According to Mosley, the frequency of the spectral line of X-rays is directly proportional to the:

(a) Mass number (b) Atomic number

(c) Atomic mass (d) Square of the atomic number

6) We can't measure the waves associated with macroscopic bodies because they have very:

(a) Small wave number (b) Small wave length (c) Large wave length (d) All

7) The maximum probability of finding the electron in H-atom is at the distance of _____ from the nucleus.

(a) 0.059 nm (b) 0.057 nm (c) 0.054 nm (d) 0. 053 nm

8) The letter 'p' in azimuthal quantum number stands for:

(a) Prominent (b) Pear-shaped (c) Prime (d) Principal

9) The idea of self-rotation of electron was given by?

(a) Bohr (b) Schrodinger

(c) Heisenberg (d) Ulenbech and Goldsmith

Maximum Marks: 16 **(SUBJECTIVE TYPE)** Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: **(12)**

i. Give importance of Mosley's law.

ii. How was dual nature of electron verified?

iii. Differentiate between orbit and orbital.

iv. What are degenerate orbitals?

v. Draw the shapes of p-orbitals.

vi. What is Pauli's exclusion principle?

SECTION-II

NOTE: Attempt All Questions: **(04)**

Q.3: What are quantum numbers? Explain principal quantum number.