

CHEMISTRY-11	Chapter#04-Seond Half (4.4-4.8) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Ionic solids are characterized by:**
 - Low melting points
 - High vapour pressures
 - Good conductivity in solid state
 - Solubility in polar solvents
- Dipole-dipole forces are present among:**
 - Molecules of Iodine
 - Chloroform molecules
 - Atoms of Neon in gaseous state
 - CCl_4 molecules
- When $a=b \neq c$ and $\alpha = \beta = 90^\circ, \gamma = 120^\circ$ then system is:**
 - Cubic
 - Triclinic
 - Hexagonal
 - Monoclinic
- The carbon atom in diamond is _____ hybridized.**
 - sp^3
 - sp^2
 - sp
 - dsp^2
- Which one is not amorphous solid?**
 - Glass
 - Plastic
 - Glue
 - Sugar
- The shape in which a crystal usually grows is called its:**
 - Unit cell
 - Symmetry
 - Habit
 - Unit cell dimension
- K_2SO_4 is isomorphic to:**
 - $ZnSO_4$
 - K_2CrO_4
 - $K_2Cr_2O_7$
 - $NiSO_4$
- The unit cell angle present between b and c is called:**
 - α
 - β
 - γ
 - Φ
- In gaseous state, the I-I bond length is:**
 - 271.5 pm
 - 266.6 pm
 - 154 pm
 - None

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

- Q.2: Give brief answers to the following questions: (12)**
- Amorphous solids like glass are also called super cooled liquids, explain.
 - Define symmetry. Name any two symmetry elements.
 - Transition temperature is shown by elements having allotropic forms and by compounds showing polymorphism.
 - Give angles and lengths of axis of monoclinic system.
 - Define lattice energy. Give example.
 - Why does the electrical conductivity of metal decrease with increase in temperature?

SECTION-II

NOTE: Attempt All Questions:

(04)

- Q.3: What are ionic solids? Give their properties.**