

CHEMISTRY-11	Chapter#03 - Second Half (3.7 to 3.11) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The deviation of gas from ideal behavior is maximum at:
 - 10°C and 5.0 atm
 - 10°C and 2.0 atm
 - 100°C and 2.0 atm
 - 0°C and 2.0 atm
- Kinetic molecular theory was put forward by:
 - Bernoulli
 - Maxwell
 - Boltzmann
 - Clausius
- The pressure of the gas is due to the collisions of the gas molecules with:
 - One another
 - The walls of the container
 - Both 'a' & 'b'
 - None
- $\sqrt{\frac{3RT}{M}}$ is called:
 - Average velocity
 - Most probable velocity
 - Mean square velocity
 - Root mean square velocity
- The coldest possible temperature in the universe is:
 - 0K
 - 0°C
 - 0°F
 - All
- The critical temperature of gas depends upon:
 - Molecular size
 - Intermolecular forces
 - Molecular shape
 - All
- Sudden expansion of gases produces:
 - Cooling effect
 - Heating effect
 - Stark effect
 - Zeeman effect
- The excluded volume 'b' is _____ times the actual volume of the gas molecules.
 - 2
 - 3
 - 4
 - 5
- Plasma constitutes about _____ of the visible universe.
 - 90%
 - 95%
 - 98%
 - More than 99%

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- Define mean square velocity and root mean square velocity.
- How does heat flow from hotter to colder body?
- Define and explain Joule Thomson effect.
- Why do gases deviate from ideal behaviour at high pressure and low temperature?
- What are the units of 'a' and 'b'?
- What are the characteristics of plasma?

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: Explain Boyle's law and Avogadro's law on the basis of KMT.