

CHEMISTRY-11	Chapter#06-First Half (6.1.0 - 6.4.3) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 10

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The radius of Na^+ ion is:
 (a) 92 pm (b) 93 pm (c) 94 pm (d) 95 pm
- Ionic, covalent and co-ordinate covalent bonds are present in:
 (a) SO_2 (b) NH_4Cl (c) C_2H_2 (d) H_2O
- Which of the following has bond angle of 120° ?
 (a) BeCl_2 (b) BF_3 (c) CH_4 (d) NH_3
- Which noble gas forms stable compounds?
 (a) Ne (b) Ar (c) Xe (d) Kr
- Which one is the correct order of the radii of H-atoms?
 (a) Covalent > atomic > anionic > cationic (b) Atomic > anionic > covalent > cationic
 (c) Anionic > covalent > atomic > cationic (d) Anionic > atomic > covalent > anionic
- Ionization energy generally _____ along the period, while _____ down the group.
 (a) Increases, decreases (b) Remains same, decreases
 (c) Decreases, increase (d) Increases, remains same
- Metalloids have _____ I.E. values.
 (a) High (b) Low (c) Intermediate (d) Very small
- The least electronegative element in the periodic table is:
 (a) H (b) Li (c) Cs (d) I
- Ionic character of NaCl:
 (a) 72% (b) 92% (c) 100% (d) 50%
- There is no bond in chemistry which has _____ ionic character.
 (a) 100% (b) 99% (c) 95% (d) All

Maximum Marks: 20

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

- Q.2: Give brief answers to the following questions:** (12)
- Bond distance is the compromise distance between two atoms. Explain.
 - Define ionic and covalent radii.
 - What are the factors which affect electron affinity?
 - Define covalent bond. Give two examples.
 - Why does lone pair of electrons occupy more space than bond pairs?
 - Why CO_2 has linear shape?

SECTION-II

NOTE: Attempt All Questions:

(08)

- Q.3: What is ionization energy. Discuss variation of ionization energy in periodic table. Also explain how ionization energy is an index to metallic character?**
- Q.4: Explain the structures of NH_3 and H_2O on the basis of VSEPR theory.**