

<b>CHEMISTRY-11</b>	<b>Chapter#06-First Half (6.1.0 - 6.4.3) Test-3</b>		
	Name:	Class:	ID:
Date: / /	<b>Marks Total: 30</b>	<b>Marks Obtained:</b>	
Time Allowed: 50 Min.			

Maximum Marks: 10

**(OBJECTIVE TYPE)**

Time Allowed: 10 Min.

**NOTE:** Tick The Correct Option:

- Which compound does not obey octet rule?  
(a)  $\text{NH}_3$  (b)  $\text{BCl}_3$  (c)  $\text{H}_2\text{O}$  (d)  $\text{CH}_4$
- Which element has highest ionization potential?  
(a) Li (b) Be (c) B (d) C
- The molecule which cannot form co-ordinate covalent bond with  $\text{H}^+$  ion is:  
(a)  $\text{NH}_3$  (b)  $\text{H}_2\text{O}$  (c)  $\text{PH}_3$  (d)  $\text{CH}_4$
- The molecular shape of  $\text{SO}_3$  is:  
(a) Triangular planar (b) Tetrahedral (c) Pyramidal (d) Linear
- When two H-atoms reach at a distance of 75.4 pm, their P.E. becomes:  
(a) Maximum (b) Minimum (c) Moderate (d) None
- Which atom will be the most stable?  
(a) With low I.E. & low E.A. (b) With high I.E. & low E.A.  
(c) With high I.E. & high E.A. (d) With low I.E. & high E.A.
- If E.N. difference between two bonded atoms is 1.7, the bond will be:  
(a) Predominantly covalent (b) 50% covalent and 50% ionic  
(c) Predominantly ionic (d) Non-polar
- Al atom loses three electrons to gain the configuration of:  
(a) He (b) Ne (c) Ar (d) Kr
- Coordinate covalent bond is also called:  
(a) Electrovalent bond (b) Electron pair bond  
(c) Dative covalent bond (d) All
- According to VSEPR theory,  $\text{H}_2\text{O}$  belongs to:  
(a)  $\text{AB}_2$  Type (b)  $\text{AB}_3$  Type (c)  $\text{AB}_4$  Type (d) None

Maximum Marks: 20

**(SUBJECTIVE TYPE)**

Time Allowed: 40 Min.

**SECTION-I**

**Q.2: Give brief answers to the following questions:**

**(12)**

- What is octet rule? Give two examples of the compounds that do not obey this rule.
- Define shielding effect?
- Name the factors which control ionization energy.
- Why is first electron affinity usually negative, while second electron affinity is always positive?
- No bond in chemistry is 100% ionic. Explain.
- The bond angle in  $\text{NF}_3$  is smaller than in  $\text{NH}_3$ . Justify.

**SECTION-II**

**NOTE:** Attempt All Questions:

**(08)**

**Q.3: Define the term electronegativity. Discuss its variation in the periodic table.**

**Q.4: Write main postulates of VSEPR theory.**